

PVT Property Measurements for Ethyl Propionate, Ethyl Butyrate, and Ethyl Pentanoate Esters from (298 to 393) K and up to 35 MPa[†]

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The results of PVT measurements of the liquid phase within the temperature range (298 to 393) K and up to 35 MPa are presented for some aliphatic esters. Measurements were made by means of a vibrating tube densimeter, model DMA 512P from Anton Parr. The calibration of the densimeter was performed with water and *n*-heptane as reference fluids. The experimental PVT data have been correlated by a Tait equation. This equation gives good results when used to predict the density of the esters using the method proposed by Thomson et al. Isothermal compressibilities, isobaric expansivities, thermal pressure coefficients, and changes in the isobaric heat capacity have been calculated from the Tait equation.

Introduction

Thermophysical properties of aliphatic esters are particularly important in the pharmaceutical, food, and flavor industries. In our research group, several studies of density, viscosity, and interfacial tensions have been made for several esters at room temperature and atmospheric pressure^{1–5} including the ethyl propionate and the ethyl butyrate. The densities of these esters have been recently investigated by other authors^{6–8} at atmospheric pressure. Detailed research investigations of these properties for pure esters, in large temperature and pressure ranges, are scarce in the literature. Recently, we have extended the pressure domain of density measurements, and we have determined reliable density data for some alkyl acetates (ethyl, propyl, *n*-butyl, and *n*-pentyl acetates) in the temperature range (298.15 to 393.15) K and up to 35 MPa.¹ In this study, new reliable density data for ethyl propionate, ethyl butyrate, and ethyl pentanoate from (298 to 393) K and up to 35 MPa are presented. For ethyl pentanoate, the reported data are very limited and are reported only at atmospheric pressure.^{10,11} As far as we know, the available literature density data for ethyl propionate and ethyl butyrate in the high-pressure domain are the measurements due to Malhotra and Woolf⁹ in the temperature range (278 to 338) K and up to 380 MPa. In this work, we have extended the temperature up to 393 K.

The density data obtained in this work have been correlated with the Tait equation using as a reference the saturated liquid equilibrium curve. From this equation, isothermal compressibilities, isobaric expansivities, thermal coefficients, and changes in isobaric heat capacity have been calculated and reported as Supporting Information.

The method of Thomson et al.¹² was applied to the prediction of the density in the pressure and temperature ranges considered since it is a model based on the Tait equation.

Experimental Section

Materials. All the materials used are proanalysis grade. Tridistilled water was used. *N*-Heptane was from Laboratory-

Scan with a stated purity higher than 99 (mass %). The esters (ethyl propionate, butyrate, and pentanoate) were obtained from Acros Organics with purities of > 99.0 mass % and have been used without further purification. The density data at different temperatures and at atmospheric pressure obtained in this work are compared with literature values in Table 1.

Experimental Procedure. Experimental densities were measured using an Anton Paar DMA 60 digital vibrating tube densimeter, with a DMA 512P measuring cell in the temperature range (293.15 to 393.15) K and pressure range (0.10 to 35.00) MPa. Figure 1 shows the installation of the DMA 512P cell and the peripheral equipment used. The temperature in the vibrating tube cell was measured with a platinum resistance probe which has a temperature uncertainty of ± 0.01 K coupled with a GW Instek Dual Display Digital Multimeter GDM-845. A Julabo P-5 thermostatic bath with silicone oil as circulating fluid was used in the thermostat circuit of the measuring cell which was held constant to ± 0.01 K.

The required pressure was generated and controlled with a Pressure Generator model 50-6-15, High Pressure Equipment Co., using acetone as the hydraulic fluid. The diameter of the metallic tube is $1.59 \cdot 10^{-3}$ m, and the buffer is more than 1 m in length which guarantees the inexistence of diffusion of the hydraulic liquid in the liquid contained in the cell of the densimeter. Pressures were measured with a pressure transducer (Wika Transmitter S-10, WIKA Alexander Wiegand GmbH & Co.) with a maximum uncertainty of ± 0.03 MPa. An NI PCI-6220 data acquisition board (DAQ) from National Instruments (NI) was used for the real time registration of values of period, temperature, and pressure. For this task, a Labview application was developed. Modules of temperature (NI SCC-FT01) and pressure (NI SCC-CI20) were installed into a NI SC-2345 carrier and connected to the DAQ board.

The calibration of the vibrating tube densimeter was described in a previous work.¹ The equation proposed by Niesen¹⁴ which has a solid theoretical basis as discussed by Holcom and Outcalt¹⁵ was used for that purpose. Using PVT data of water and *n*-heptane as reference fluids, the standard deviation of the fitted calibration curve was $1 \text{ kg} \cdot \text{m}^{-3}$. As shown in our previous work,¹ the influence of the viscosity on the density is insignificant.

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Table 1. Densities of Pure Esters at Different Temperatures and at Atmospheric Pressure, Obtained in this Work and from the Literature

ester	T/K	$\rho/\text{kg}\cdot\text{m}^{-3}$		
		this work	literature	
ethyl propionate	298.15	884.3	884.359 ^a ; 884.16 ^b ; 884.0 ^c	
	303.15	878.5	878.87 ^d ; 879.14 ^e ; 879.03 ^b	
	308.15	872.7	873.33 ^e	
	313.15	866.8	867.094 ^a	
	318.15	860.9	861.60 ^e	
	323.15	854.9	855.456 ^a	
	328.15	849.2	849.85 ^e	
	333.15	843.1	843.96 ^e	
	338.15	836.9	837.671 ^a	
	343.15	831.1		
	348.15	824.9		
	353.15	818.7		
	363.15	803.5		
	373.15	790.9		
	ethyl butyrate	298.15	873.3	873.668 ^a ; 873.47 ^f ; 873.94 ^e ; 874.32 ^e
		303.15	868.1	868.21 ^g ; 868.83 ^g ;
		308.15	862.6	863.40 ^e
313.15		857.3	857.765 ^a ; 857.98 ^e	
318.15		851.9	852.46 ^e	
323.15		846.7	847.031 ^a ; 847.03 ^e	
328.15		841.4	841.54 ^e	
333.15		835.8		
338.15		830.3	830.781 ^a	
343.15		824.8		
348.15		819.1		
353.15		813.5		
363.15		799.9		
373.15		788.5		
383.15		776.9		
393.15		765.1		
ethyl pentanoate		298.15	870.7	869.39 ^h
	303.15	865.5	866.90 ⁱ	
	308.15	860.3		
	313.15	855.0	856.81 ⁱ	
	318.15	849.8		
	323.15	844.4		
	328.15	839.1		
	333.15	833.7		
	338.15	828.4		
	343.15	823.0		
	348.15	817.7		
	353.15	812.2		
	363.15	801.2		
	373.15	790.7		
	383.15	779.9		
	393.15	769.0		

^a Ref 9. ^b Ref 6. ^c Ref 13. ^d Ref 2. ^e Ref 7. ^f Ref 8. ^g Ref 4. ^h Ref 10. ⁱ Ref 11.

Results and Discussion

Density measurements were carried out in wide ranges of temperatures, (298 to 393) K, and pressures, (0.1 to 35) MPa. Particular attention was given to the subcritical data, and therefore we have measured the density in the above-mentioned temperature domain for several subcritical pressure values. The experimental data for ethyl propionate, ethyl butyrate, and ethyl pentanoate are reported in Tables 2 to 4. In Figure 2, we represent the isotherms of the density, ρ , for ethyl propionate. The expected behavior with temperature and pressure is found. The densities of the esters decrease with the increase of the alkanate chain length, and this effect is more important at low temperature. This behavior was also observed in our previous reported data for alkyl acetates¹ and was discussed on the basis that addition of $-\text{CH}_2-$ groups disrupts the local configurational order of the $-\text{COO}-$ group causing a lower molecular packing efficiency which leads to the decrease in the density.¹⁶

The density data due to Malhotra and Woolf⁹ for ethyl propionate and ethyl butyrate at temperatures of (298.14, 313.14,

323.14, and 338.13) K have been compared with our own measurements at the same temperatures and pressures. The results of comparison are shown in Figure 3. The average absolute deviation, AAD, defined as

$$\text{AAD} (\%) = \frac{\sum_{i=1}^{N_p} |(\rho_{\text{exptl}} - \rho_{\text{lit}}.) / \rho_{\text{lit}}|}{N_p} \quad (1)$$

is 0.2 % for both substances. In eq 1, ρ_{exptl} and ρ_{lit} are, respectively, the experimental density data from this work and the experimental values found by Malhotra and Woolf⁹ for the same temperature and pressure, and N_p represents the number of points.

The Tait equation

$$\rho = \frac{\rho^{\text{ref}}}{\left\{ 1 - C \ln \frac{(B + P)}{(B + P^{\text{ref}})} \right\}} \quad (2)$$

was fitted to the experimental density data in the ranges of pressures and of temperatures mentioned above. As discussed in a previous work,¹ the saturated liquid state (σ) was selected here as the reference state which means $\rho^{\text{ref}} = \rho^{\prime\prime\sigma}$ and $P^{\text{ref}} = P^\sigma$. The vapor pressure (P^σ) of the ethyl propionate and ethyl butyrate was obtained from the Wagner equation

$$\ln(P^\sigma/\text{kPa}) = \ln(P_c/\text{kPa}) + (T_c/T)(n_1\tau + n_2\tau^{1.5} + n_3\tau^3 + n_4\tau^6) \quad (3)$$

where $\tau = 1 - T/T_c$. The critical constants (T_c , P_c) and the parameters n_i are listed in Table 5. For ethyl pentanoate we have used the Antoine equation

$$\ln(P^\sigma/\text{kPa}) = n_1 - n_2/[(T/\text{K}) + n_3] \quad (4)$$

The values of the saturated liquid density for a given temperature were obtained from values of density data measured at pressures lower than the critical, using eqs 3 and 4 as explained before.¹

The density data of the saturated liquid ($\rho^{\prime\prime\sigma}$) are reported in Table 6. The Rackett equation, modified by Spencer and Danner²¹

$$\rho^{\prime\prime\sigma} = \left\{ \frac{MP_c}{RT_c} \right\} \frac{1}{Z_{\text{RA}}^{[1+(1-T_r)^{2/7}]}} \quad (5)$$

was fitted to the density data of the saturated liquid in the range (298.15 to 393.15) K. The values of the fitted parameter Z_{RA} and the associated standard deviation of the fit, σ , defined as

$$\sigma = \left[\sum_{i=1}^{N_p} (\rho_{\text{calcd}}^{\prime\prime\sigma} - \rho_{\text{exptl}}^{\prime\prime\sigma})^2 / (N_p - k) \right]^{1/2} \quad (6)$$

where $\rho_{\text{calcd}}^{\prime\prime\sigma}$ and $\rho_{\text{exptl}}^{\prime\prime\sigma}$ are, respectively, the density data from eq 5 determined as described above and k is the number of adjusted parameters ($k = 1$), are listed in Table 7. In the same table, values of density at the critical point (ρ_c) obtained from eq 5 are compared with literature ones. The accordance is good. The calculated values of the saturated liquid density from eq 5 are represented in Figure 4, and our values of $\rho_{\text{exptl}}^{\prime\prime\sigma}$ are compared with those from the literature for ethyl propionate. As for the other alkyl acetates studied,¹ the agreement is good and the saturation curves can be well predicted in all the liquid + gas equilibrium ranges using eq 5 fitted to density data in a very narrow range of temperatures.

Table 2. Experimental Density, ρ , Data for Ethyl Propionate as a Function of Temperature, T , and Pressure, P

T/K	$\rho/\text{kg}\cdot\text{m}^{-3}$ at P/MPa											
	0.10	1.00	2.00	3.00	3.34	5.00	10.00	15.00	20.00	25.00	30.00	35.00
298.15	884.3	884.8	885.7	886.6	886.9	888.7	892.9	896.9	900.7	904.4	908.9	913.6
303.15	878.5	879.0	880.0	880.9	881.2	883.0	887.4	891.6	895.6	899.4	904.0	908.9
308.15	872.7	873.2	874.2	875.2	875.5	877.4	882.0	886.3	890.4	894.3	899.2	904.2
313.15	866.8	867.4	868.4	869.4	869.8	871.7	876.5	880.9	885.1	889.3	894.4	899.5
318.15	860.9	861.5	862.6	863.7	864.1	866.1	871.0	875.6	880.0	884.1	889.1	894.3
323.15	854.9	856.0	857.1	858.1	858.4	860.3	865.4	870.2	874.7	879.1	884.3	889.7
328.15	849.2	850.2	851.3	852.5	852.9	854.7	860.0	865.0	869.6	874.0	879.4	884.9
333.15	843.1	844.2	845.4	846.6	847.0	849.0	854.6	859.7	864.5	869.2	874.8	880.6
338.15	836.9	838.2	839.5	840.7	841.1	843.1	848.8	854.4	859.4	863.6	869.9	875.8
343.15	831.1	832.2	833.5	834.8	835.2	837.4	843.4	848.9	854.0	859.1	865.0	871.0
348.15	824.9	826.0	827.4	828.8	829.3	831.5	837.7	843.4	848.8	854.0	860.1	866.3
353.15	818.7	819.9	821.3	822.7	823.2	825.6	832.1	838.0	843.5	848.9	855.3	861.6
363.15	803.5	805.1	806.9	808.7	809.3	812.0	819.6	826.5	833.0	839.1	844.8	850.2
373.15	790.9	792.6	794.5	796.4	797.1	800.0	808.2	815.6	822.5	829.0	834.9	840.6
383.15	—	780.0	782.1	784.1	784.8	787.9	796.6	804.6	811.9	818.7	825.0	830.9
393.15	—	766.4	768.7	770.8	771.5	775.0	784.5	793.1	800.8	808.0	814.7	820.9

Table 3. Experimental Density, ρ , Data for Ethyl Butyrate as a Function of Temperature, T , and Pressure, P

T/K	$\rho/\text{kg}\cdot\text{m}^{-3}$ at P/MPa											
	0.10	1.00	2.00	3.00	3.10	5.00	10.00	15.00	20.00	25.00	30.00	35.00
298.15	873.3	874.1	875.0	875.9	876.0	877.5	881.5	885.3	889.0	892.5	896.9	901.4
303.15	868.1	868.9	869.8	870.7	870.7	872.4	876.5	880.4	884.1	887.8	892.3	896.9
308.15	862.6	863.5	864.4	865.3	865.4	867.1	871.4	875.5	879.4	883.1	887.8	892.5
313.15	857.3	858.2	859.1	860.1	860.2	861.9	866.3	870.5	874.5	878.4	883.2	888.1
318.15	851.9	852.8	853.8	854.8	854.9	856.7	861.3	865.8	869.9	874.1	879.0	884.1
323.15	846.7	847.4	848.4	849.5	849.6	851.7	856.7	861.1	865.4	869.4	874.5	879.7
328.15	841.4	842.0	843.3	844.5	844.6	846.6	851.5	856.2	860.5	864.8	870.0	875.3
333.15	835.8	837.0	838.1	839.2	839.3	841.2	846.3	851.1	855.6	860.1	865.4	870.8
338.15	830.3	831.5	832.6	833.8	833.9	835.9	841.2	846.2	850.8	855.5	861.0	866.5
343.15	824.8	825.9	827.1	828.3	828.4	830.7	836.2	841.4	846.1	850.1	855.7	861.8
348.15	819.1	820.3	821.6	822.9	823.0	825.1	830.9	836.2	841.1	846.0	851.8	857.7
353.15	813.5	814.7	816.1	817.4	817.5	819.8	825.7	831.2	836.3	841.4	847.4	853.4
363.15	799.9	801.4	803.0	804.5	804.7	807.9	814.4	820.7	826.7	832.4	837.7	842.7
373.15	788.5	790.1	791.7	793.4	793.6	796.8	804.0	810.8	817.2	823.1	828.7	834.0
383.15	776.9	778.6	780.4	782.2	782.4	785.5	793.5	800.8	807.5	813.8	819.6	825.2
393.15	765.1	766.9	768.8	770.8	771.0	774.5	783.1	790.9	798.0	804.6	810.7	816.6

Table 4. Experimental Density, ρ , Data for Ethyl Pentanoate as a Function of Temperature, T , and Pressure, P

T/K	$\rho/\text{kg}\cdot\text{m}^{-3}$ at P/MPa											
	0.10	1.00	2.00	2.78	3.00	5.00	10.00	15.00	20.00	25.00	30.00	35.00
298.15	870.7	871.5	872.4	873.0	873.2	875.1	879.4	883.4	887.3	891.0	894.6	898.1
303.15	865.5	866.4	867.4	868.1	868.3	870.1	874.5	878.7	882.6	886.5	890.2	893.8
308.15	860.3	861.2	862.1	862.9	863.1	865.0	869.5	873.8	877.8	881.8	885.6	889.2
313.15	855.0	855.9	856.9	857.7	857.9	859.9	864.5	868.9	873.0	877.0	880.9	884.6
318.15	849.8	850.6	851.7	852.4	852.7	854.6	859.4	863.9	868.2	872.3	876.3	880.1
323.15	844.4	845.3	846.4	847.2	847.4	849.4	854.3	859.0	863.4	867.6	871.7	875.5
328.15	839.1	840.0	841.1	842.0	842.3	844.4	849.4	854.2	858.7	863.0	867.1	871.1
333.15	833.7	834.7	836.0	836.9	837.1	839.3	844.5	849.5	854.1	858.6	862.9	867.0
338.15	828.4	829.5	830.7	831.5	831.8	834.0	839.4	844.5	849.3	853.9	858.2	862.4
343.15	823.0	824.1	825.3	826.2	826.5	828.8	834.4	839.6	844.5	849.2	853.7	858.0
348.15	817.7	818.7	820.0	821.0	821.3	823.7	829.4	834.8	839.9	844.7	849.3	853.6
353.15	812.2	813.3	814.6	815.6	815.9	818.4	824.3	829.9	835.1	840.0	844.7	849.1
363.15	801.2	802.5	803.9	805.0	805.5	808.1	814.6	820.5	826.0	831.2	836.0	840.8
373.15	790.7	792.0	793.5	794.7	795.0	797.9	804.8	810.9	816.7	822.2	827.3	832.3
383.15	779.9	781.3	783.0	784.2	784.5	787.6	795.0	801.5	807.6	813.4	818.9	824.0
393.15	769.0	770.5	772.3	773.6	774.1	777.4	785.1	791.6	797.7	803.5	809.0	814.0

In eq 2, the coefficient C was treated as constant for each substance while the parameter B was considered temperature dependent following a function proposed by Kumagai and Iwasaki.²³ They have found that B is proportional to the reciprocal of temperature. This dependence is given by

$$B = b_1 + \frac{b_2}{T} \quad (7)$$

where b_1 and b_2 are parameters obtained from the fitting of eq 2 to PVT data. In Table 8, we list the parameters of the Tait equation, the average absolute deviation (AAD), and the standard deviation of the fittings (σ) for the esters studied.

Table 5. Critical Constants (T_c, P_c) and Parameters of the Wagner Equation, Equation 3^c and the Antoine Equation, Equation 4^b

ester	T_c	P_c	parameters of eqs 3 and 4			
	K	kPa	n_1	n_2	n_3	n_4
ethyl propionate	546.1 ^c	3360 ^c	-7.76801	1.19545	-3.81174	-3.81174
ethyl butyrate	569.0 ^c	2950 ^c	-7.93351	1.22092	-3.89295	-3.89295
ethyl pentanoate	593.3 ^d	2780 ^e	15.3818	4074.13	-40.59	

^a The parameters are from ref 17. ^b Parameters from ref 18. ^c Ref 17. ^d Ref 19. ^e Ref 20.

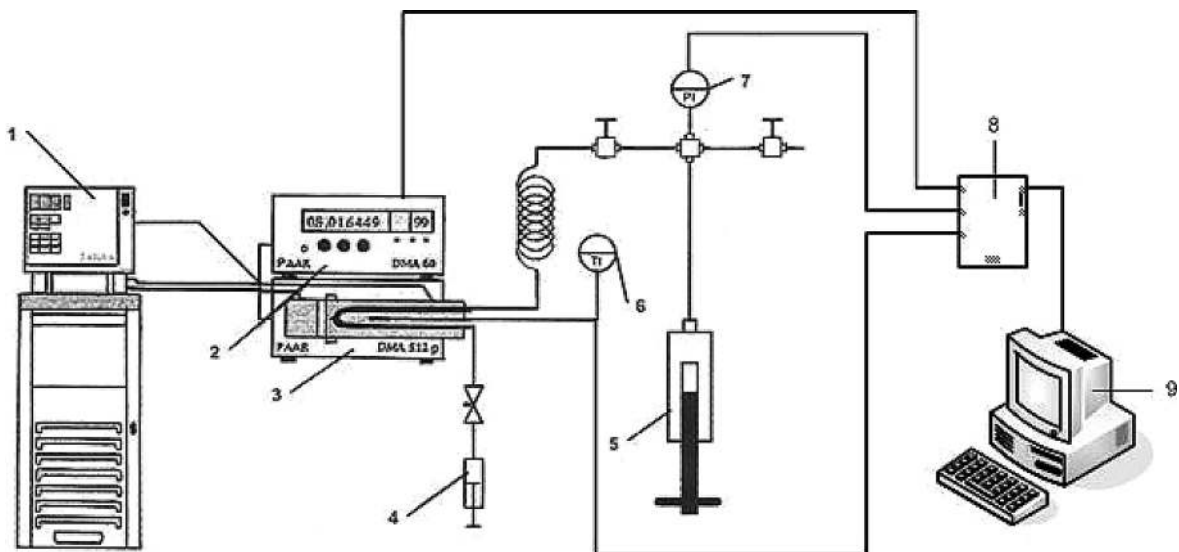


Figure 1. Experimental setup for the measurement of liquid densities at high pressures: 1, Julabo FP-50 thermostatic bath; 2, DMA 60 (Anton Paar) device for measuring the period of oscillation; 3, measuring cell DMA 512P (Anton Paar); 4, syringe for sample introduction; 5, pressure generator model HIP 50-6-15; 6, PT probe; 7, pressure transducer WIKA, S-10; 8, NI PCI-8220 data acquisition board; 9, PC.

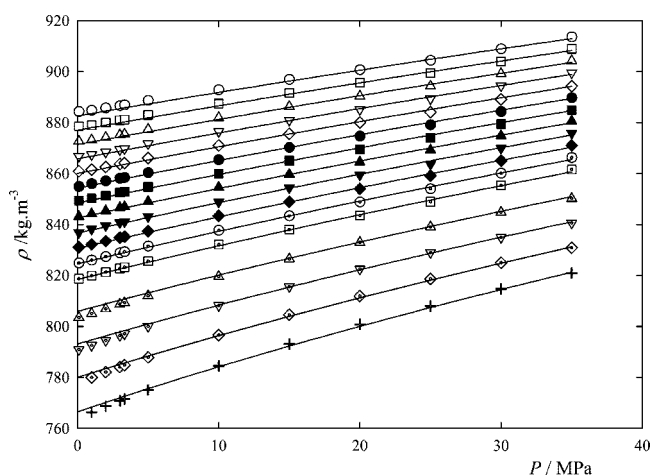


Figure 2. Isotherms of the density (ρ) for ethyl propionate. The symbols refer to the experimental data: from \circ , 298.15 K to \square , 353.15 K with increments of 5 K, and from \square , 353.15 K to $+$, 393.15 K with increments of 10 K. The curves are calculated with eq 2.

In Figure 5, the percentage density deviation plots between the calculated values with eq 2 and experimental data are shown. The absolute percentage deviations are less than 0.2 % in the majority part of the temperature and pressure domains. The AAD is of the order of 0.1 % for all the substances (see Table 8).

Some mechanical coefficients were obtained from the Tait equation. For the isothermal compressibility (k_T) defined as

$$k_T = -\frac{1}{V_m} \left(\frac{\partial V_m}{\partial P} \right)_T = \frac{1}{\rho} \left(\frac{\partial \rho}{\partial P} \right)_T = \left(\frac{\partial \ln \rho}{\partial P} \right)_T \quad (8)$$

the final expression is

$$k_T = \left(\frac{C}{B+P} \right) \left(\frac{\rho}{\rho^{\prime\sigma}} \right) \quad (9)$$

The isothermal compressibilities of the esters are listed as Supporting Information, and the values for ethyl propionate are shown in Figure 6. For the other alkyl esters, the representations are similar.

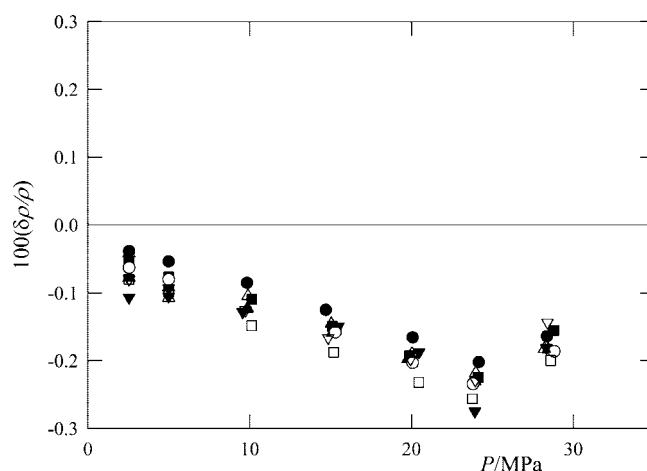


Figure 3. Percentage deviations between experimental data of this work and that from Malhotra and Woolf⁹ at the temperatures of \bullet , \circ , 298.14 K; \blacksquare , \square , 313.14 K; \blacktriangle , \triangle , 323.14 K; \blacktriangledown , \triangledown , 338.13 K. The full symbols refer to the ethyl propionate and the open symbols to ethyl butyrate. The quantity $100(\delta\rho/\rho)$ means $100(\rho_{\text{eq 2}} - \rho_{\text{lit}})/\rho_{\text{lit}}$.

For the isobaric expansivity (α_P)

$$\alpha_P = \frac{1}{V_m} \left(\frac{\partial V_m}{\partial T} \right)_P = -\frac{1}{\rho} \left(\frac{\partial \rho}{\partial T} \right)_P = -\left(\frac{\partial \ln \rho}{\partial T} \right)_P \quad (10)$$

the following expression is derived from the Tait and the Rackett equations

$$\alpha_P = -\frac{2}{7} \frac{\ln(Z_{RA})}{T_c(1-T_r)^{5/7}} - C \left\{ \frac{\frac{dB}{dT}(P^\sigma - P) - \gamma^\sigma(B+P)}{\left[1 - C \ln \left(\frac{B+P}{B+P^\sigma} \right) \right] (B+P^\sigma)(B+P)} \right\} \quad (11)$$

where $dB/dT = -b_2/T^2$ and $\gamma^\sigma = dP^\sigma/dT$ is the slope of the vapor pressure curve at the given temperature T . The values of α_P are plotted in Figure 7 for ethyl butyrate. Values of this mechanical coefficient for the other esters are given as Supporting Information.

Table 6. Density of the Saturated Liquid, $\rho^{\ell,\sigma}$

T	$\rho^{\ell,\sigma}$	T	$\rho^{\ell,\sigma}$	T	$\rho^{\ell,\sigma}$	T	$\rho^{\ell,\sigma}$
K	$\text{kg}\cdot\text{m}^{-3}$	K	$\text{kg}\cdot\text{m}^{-3}$	K	$\text{kg}\cdot\text{m}^{-3}$	K	$\text{kg}\cdot\text{m}^{-3}$
Ethyl Propionate							
298.15	884.1	318.15	860.7	338.15	836.9	363.15	803.5
303.15	878.3	323.15	854.9	343.15	831.0	373.15	790.9
308.15	872.5	328.15	849.1	348.15	824.8	383.15	778.3
313.15	866.6	333.15	843.0	353.15	818.6	393.15	764.7
Ethyl Butyrate							
298.15	873.2	318.15	851.8	338.15	830.2	363.15	799.8
303.15	868.0	323.15	846.5	343.15	824.7	373.15	788.4
308.15	862.5	328.15	841.1	348.15	819.0	383.15	776.9
313.15	857.2	333.15	835.8	353.15	813.4	393.15	765.1
Ethyl Pentanoate							
298.15	870.6	318.15	849.7	338.15	828.3	363.15	801.1
303.15	865.4	323.15	844.3	343.15	822.9	373.15	790.6
308.15	860.2	328.15	839.0	348.15	817.5	383.15	779.8
313.15	855.0	333.15	833.6	353.15	812.1	393.15	768.9

Table 7. Coefficient Z_{RA} in the Rackett Equation, Equation 5, Standard Deviation of the Fit, σ , and Average Absolute Deviation, AAD

ester	Z_{RA}	$\sigma/\text{kg}\cdot\text{m}^{-3}$	AAD %	$\rho_c/\text{kg}\cdot\text{m}^{-3}$	eq 5	lit.
ethyl propionate	0.2549 ± 0.0001	1.4	0.13	296.5	296.0 ^a	
ethyl butyrate	0.2527 ± 0.0000	1.1	0.10	286.6	275.9 ^a	
ethyl pentanoate	0.2573 ± 0.0001	2.1	0.21	285.2	289.9 ^a	

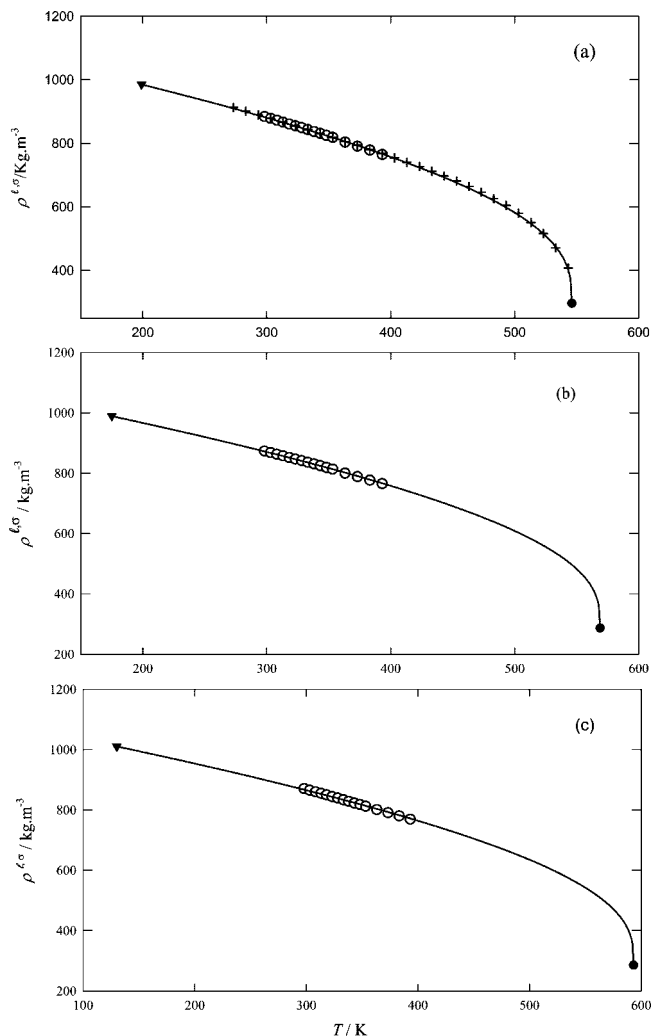
^a Ref 20.

The thermal pressure coefficient (γ_V) was calculated according to

$$\gamma_V = \left(\frac{\partial P}{\partial T} \right)_\rho = \frac{\alpha_P}{k_T} \quad (12)$$

and the values for this property are plotted in Figure 8 for ethyl pentanoate. The thermal pressure coefficient data for ethyl propionate and ethyl butyrate are also given as Supporting Information. We have performed a statistical study to determine the correct number of significant figures of the mechanical coefficients. We have applied the propagation law of errors at eqs 9, 11, and 12 taking into account the uncertainties in the parameters Z_{RA} , C , b_1 , and b_2 and also in the temperature and pressure. We have concluded that the calculated k_T must be presented with no more than one decimal place, and the α_P can be written with two or one decimals depending on the (T , P) domain considered. Obviously γ_V must be presented with only one decimal place. The values of α_P are similar in magnitude for all the esters in the tested ranges of temperature and pressure. The same is true for k_T . The dependence on P and T of these mechanical coefficients follows the expected behavior. The coefficients α_P and k_T show an increase with the temperature which is more significant at lower pressures. At high pressures, the value of α_P is practically constant, independent of temperature. At fixed pressure, the value of the thermal pressure coefficient decreases with the increasing temperature, and an isothermal increase with pressure in this property is observed, especially at values of high temperature isotherms. From the tables given in the Supporting Information, it is observed that the effect of adding $-\text{CH}_2-$ units to the alkanate group of the ester molecule has negligible changes on the mechanical coefficients.

The α_P and k_T data from this work have been compared with the values reported by Malhotra and Woolf⁹ for ethyl propionate and ethyl butyrate. The absolute percentage deviation for both coefficients is between (1 and 3) %, and therefore the agreement can be considered good because these coefficients involve the derivatives of density which are highly sensible to minor changes in the density.

**Figure 4.** Density of the saturated liquid of the alkyl esters ($\rho^{\ell,\sigma}$) as a function of temperature. Full curves are calculated from eq 5 with the value Z_{RA} from this work: \blacktriangledown , triple point; \bullet , critical point. (a) Ethyl propionate: $+$, ref 22; (b) ethyl butyrate; (c) ethyl pentanoate. \circ , This work.**Table 8. Coefficients of Equation 2 and Equation 7, Standard Deviation of the Fit, σ , and Average Absolute Deviation, AAD**

ester	C	b_1	b_2	σ	AAD
		MPa	MPa·K	$\text{kg}\cdot\text{m}^{-3}$	%
ethyl propionate	0.138	-170.04	88936.6	0.8	0.07
ethyl butyrate	0.140	-171.8	92378.0	0.7	0.07
ethyl pentanoate	0.103	-48.123	40984.3	1.4	0.14

Taking into account the influence of pressure on the isobaric molar heat capacity ($C_{P,m}$) at constant temperature

$$\left(\frac{\partial C_{P,m}}{\partial P} \right)_T = -T \left(\frac{\partial^2 V_m}{\partial T^2} \right)_P \quad (13)$$

the following equation is obtained for the change in the isobaric molar heat capacity due to the increase in the pressure from 0.1 MPa to P

$$\Delta C_{P,m}(T, P) = C_{P,m}(T, P) - C_{P,m}(T, P = 0.1 \text{ MPa}) = -T \int_{0.1}^P \left\{ \frac{M}{\rho} \left[\left(\frac{\partial \alpha_P}{\partial T} \right)_P + (\alpha_P)^2 \right] \right\} dP \quad (14)$$

where M is the molar mass. The quantities in the integrand of eq 14 were calculated from the Tait equation. The values of $\Delta C_{P,m}$ are given also in the Supporting Information and can be

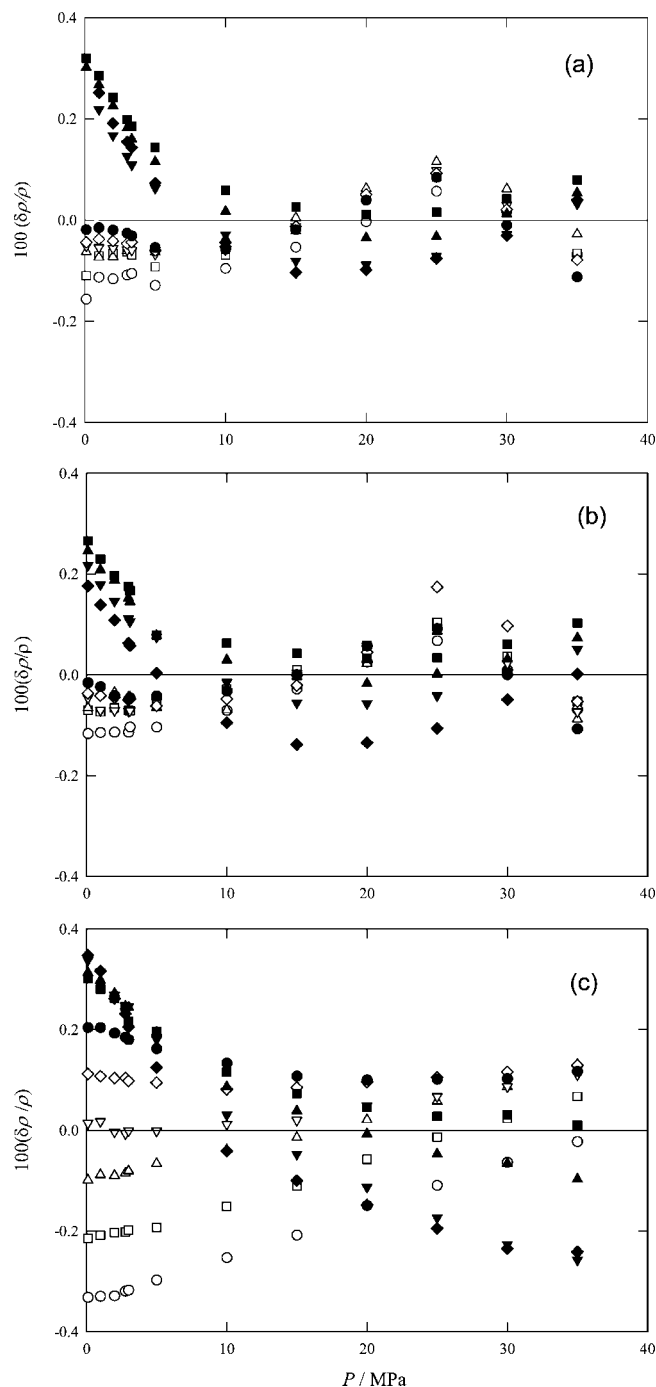


Figure 5. Percentage density deviations between the values obtained with eq 2 and experimental data: (a) ethyl propionate; (b) ethyl butyrate; (c) ethyl pentanoate. Experimental data at the temperatures of ○, 303.15 K; □, 313.15 K; △, 323.15 K; ▽, 333.15 K; ◇, 343.15 K; ●, 353.15 K; ■, 363.15 K; ▲, 373.15 K; ▼, 383.15 K; ◆, 393.15 K. The quantity $100(\delta\rho/\rho)$ means $100(\rho_{\text{eq } 2} - \rho_{\text{expt}})/\rho_{\text{expt}}$.

estimated to have an error of $(\pm 1 \text{ to } 2) \text{ J}\cdot\text{mol}^{-1}\cdot\text{K}^{-1}$. The most appreciable changes in $C_{P,m}$ occur at high pressure and high temperature. However, in the range of pressures covered by this work, the $\Delta C_{P,m}$ is small compared with the value of $C_{P,m}(T, P = 0.1 \text{ MPa})$. Following the data compiled by Zábbranský et al.²⁴ for ethyl propionate, the highest temperature available for comparison is 350 K. At this temperature and atmospheric pressure, $C_{P,m} = 211 \text{ J}\cdot\text{mol}^{-1}\cdot\text{K}^{-1}$ and $\Delta C_{P,m}(350 \text{ K}, 35 \text{ MPa}) = -8 \text{ J}\cdot\text{mol}^{-1}\cdot\text{K}^{-1}$. For ethyl butyrate, $C_{P,m}(350 \text{ K}, 0.1 \text{ MPa}) = 248 \text{ J}\cdot\text{mol}^{-1}\cdot\text{K}^{-1}$ and $\Delta C_{P,m}(350 \text{ K}, 35 \text{ MPa}) = -7 \text{ J}\cdot\text{mol}^{-1}\cdot\text{K}^{-1}$.

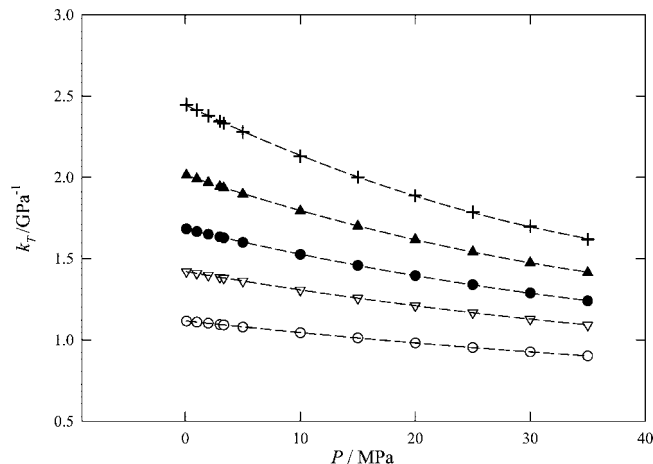


Figure 6. Isothermal compressibility of ethyl propionate at the temperatures: ○, 303.15 K; ▽, 333.15 K; ●, 353.15 K; ▲, 373.15 K; +, 393.15 K. The curves were fitted to data.

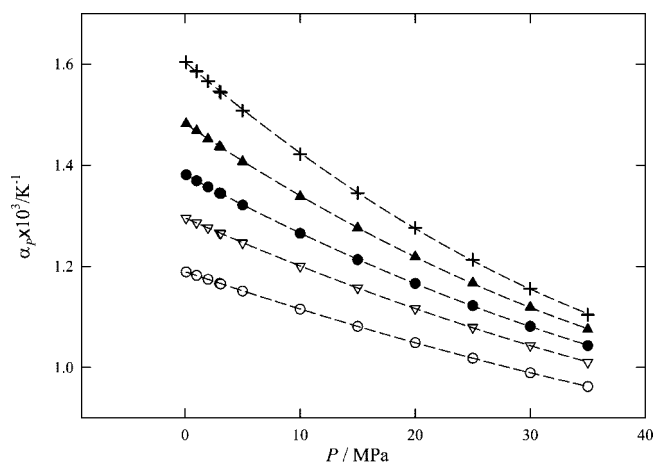


Figure 7. Isobaric expansivity of ethyl butyrate at the temperatures: ○, 303.15 K; ▽, 333.15 K; ●, 353.15 K; ▲, 373.15 K; +, 393.15 K. The curves were fitted to data.

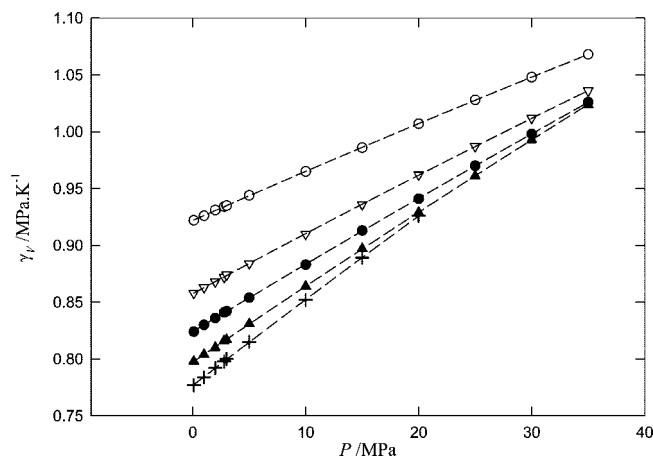


Figure 8. Thermal pressure of ethyl pentanoate at the temperatures: ○, 303.15 K; ▽, 333.15 K; ●, 353.15 K; ▲, 373.15 K; +, 393.15 K. The curves were fitted to data.

The Tait equation was also used to predict the density of the ethyl alkanoates studied here. For this purpose, the method due to Thomson et al.¹² was applied. In this method, the parameters B and C of eq 1 are predicted by the following equations

$$B/P_c = -1.0 + a(1 - T_r)^{1/3} + b(1 - T_r)^{2/3} + d(1 - T_r) + e(1 - T_r)^{4/3} \quad (15)$$

$$C = j + k\omega \quad (16)$$

$$e = \exp(f + g\omega + h\omega^2) \quad (17)$$

were ω is the Pitzer acentric factor and $a = -9.070217$, $b = 62.45326$, $d = -135.1102$, $f = 4.79594$, $g = 0.250047$, $h = 1.14188$, $j = 0.0861488$, and $k = 0.0344483$. The necessary vapor pressure and saturated liquid densities to be used in eq 2 were provided by eqs 3 to 5. The AAD between experimental and predicted values of density is 0.6 %.

Conclusions

In this work, experimental density data of ethyl alkanooates (propionate, butyrate, and pentanoate) over the pressure range (0.1 to 35) MPa and the temperature range (298.15 to 393.15) K are presented. We have extended the temperature domains of the experimental measurements for ethyl propionate and ethyl butyrate considering the available literature data. Our measurements are in close agreement in the (P , T) domain covered by other authors. As far as we know, density data for the ethyl pentanoate at wide ranges of temperature and pressure are original. We have found that the density of the esters is essentially constant and decreases only slightly with the increase of the alkyl chain length. This behavior was already found in alkyl acetates investigated in a previous work.

The Tait equation was fitted to the density data using simultaneously the pressure and the temperature which is a more direct and fast methodology than to fit the equation to isothermal (ρ , P) data, as some authors do. With this procedure, the mechanical coefficients α_p , k_T , and γ_v and the change in the isobaric molar heat capacity ($\Delta C_{p,m}$), from $P = 0.1$ MPa to P at fixed T were calculated in wide (P , T) ranges. The values of these PVT properties show a good agreement with those from other authors. The calculated values of the change in the isobaric heat capacity ($\Delta C_{p,m}$) are always negative and represent less than 5 % of the values of $C_{p,m}(T, P = 0.1 \text{ MPa})$.

The method due to Thomson et al. can be applied to the prediction of density of the esters with satisfactory results.

Supporting Information Available:

An additional four tables. This material is available free of charge via the Internet at <http://pubs.acs.org>.

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