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ARTICLE TYPE

Alkyltributylphosphonium chloride ionic liquids: synthesis, physicochemical properties and crystal structure

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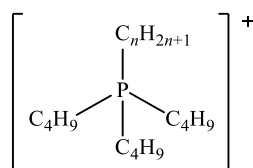
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A series of alkyltributylphosphonium chloride ionic liquids, prepared from tributylphosphine and the respective 1-chloroalkane, $C_nH_{2n+1}Cl$ (where $n = 1, 2, 3, 4, 5, 6, 7, 8, 9, 10, 12$ or 14) is reported. This work is a continuation of an extended series of tetraalkylphosphonium ionic liquids, where the focus is on the variability of n and its impact on the physical properties, such as melting points / glass transitions, thermal stability, density and viscosity. Experimental density and viscosity data were interpreted using QPSR and group contribution methods and the crystal structure of propyl(tributyl)phosphonium chloride is detailed.

Introduction

15 Tetraalkylphosphonium ionic liquids are a readily available family of ionic liquids that, in many applications, offer superior properties when compared to nitrogen-based ionic liquids.¹⁻⁶ Applications investigated in recent years include their use as extraction and synthetic solvents, electrolytes in batteries, corrosion protection, or in petrochemical applications.¹

The most attention in the literature has been given to the $[P_{66614}]^+$ cation,⁷⁻¹⁰ as this is commercially available with a range of different anions, and relatively cheap. However, there are very few systematic studies of the physicochemical and thermal properties of ionic liquids prepared with different phosphonium cations. As a detailed systematic analysis of fundamental properties of tetraalkylphosphonium ionic liquids is necessary for future application of these compounds, we report here the properties of a related series of hydrophilic ionic liquids, $[P_{444n}]Cl$, where $n = 1-8, 10, 12$ or 14 (see **Figure 1**), to complement data recently published from our group relating to the hydrophobic $[P_{888n}]Cl$.¹¹



35 **Fig. 1** The structure of the alkyltributylphosphonium cation, $[P_{444n}]^+$, where $n = 1, 2, 3, 4, 5, 6, 7, 8, 10, 12$ or 14 .

The variation of length in the variable alkyl chain has been studied with respect to its effect on the density, viscosity, melting points/glass transitions and thermal stability. Related studies based on the alkyltriethylphosphonium cations will be published shortly.^{11, 12} Although there are a few recent papers reporting

tetraalkylphosphonium salts with different anions, the most interesting ones feature anions derived from amino acids.¹³⁻¹⁵ However, reports of alkyltributylphosphonium halides date back the 1980s, when Knifton (a Texaco chemist) published a fascinating series of patents and papers in which the "molten salts" ($[P_{44416}]Br$, $[P_{4444}]I$, $[P_{4444}]Cl$ or $[P_{4444}][O_2CMe]$, with the cynosure on $[P_{4444}]Br$) were used as the basis for the hydroformylation of internal olefins in the presence of a range of ruthenium catalysts.^{16,17-26} Alkyltributylphosphonium halides were successfully applied in late 80s as biocides²⁷ or as compounds with microbiological activity for water treatment.²⁸ The series presented here has also been tested for toxicity with respect to a specific group of fungi, by Pereira's group in Portugal, and these data will be reported elsewhere.²⁹

55 Despite the impact of Knifton's studies, it was not until 1990 that tributylphosphine was produced on a large commercial scale. Since that time, not only have tetraalkylphosphonium chloride and bromide become available on the multi-ton scale, but so have many trialkylphosphines.⁷ This manuscript illuminates the versatility and practical applicability of ionic liquids derived from tributylphosphine by making available for the first time a systematic study of their physical and structural properties and contrasts them with those of related tetraalkylammonium salts.

Experimental

65 Instrumental

Autoclave. The miniclave drive manufactured by Büchi Glas Uster (Switzerland) was purchased from Ken Kimbel & Co. Ltd. (United Kingdom). The autoclave consists of a glass reactor with nominal volume 200 cm³; holder for glass reactor with stainless steel anti-splinter cage; insert ring; cover plate with 6 welded-in ¼ " Swagelok connections, welded on bmd 075 magnetic drive; O-ring seal; union nut and holding rod. The magnetic coupling

(bmd 075) is combined with the cyclone 075 (0-2500 rpm) equipped with an electric motor, into a hermetically sealed unit. The controller is equipped with a remote control input for external speed adjustment. The stationary housing of the magnetic drive and the cover plate are welded together and built as one common unit. The cover plate contains pressure relief valve, platinum resistance thermometer (Pt 100), gas inlet tube, pressure gauge and an inlet/outlet tube. The maximum temperature is 250 °C and maximum pressure is 10 bar.

Microwave. MicroSYNTH microwave used was manufactured by Milestone (Italy) as the single-mode set up with 20 cm³ volume vial placed in the TMF safety shield vessel, closed by TMF screw cap and located in the microwave cavity (43 l) at the position with the highest energy intensity allowing rapid heating. Reaction temperature control is provided by an IR sensor (max. 250 °C), monitoring the outside surface temperature of the vessel, mounted in the sidewall of the cavity, about 5 cm above the bottom and fiber-optic sensor (max. 300 °C) immersed inside of the vessel. A pneumatic pressure sensor (max. 55 bar) placed inside of the vessel monitors the vapour concentration and overall pressure. An additional sensor monitors the pressure inside of the cavity. Post-reaction cooling of the reaction mixture is achieved by a constant airflow (1.8 m³ min⁻¹) through the cavity and a stream of compressed air. The MicroSYNTH reaction system is monitored *via* an external control terminal with a touch screen display utilising the EasyControl 640 software package. The run can be controlled either by temperature, pressure or microwave output (max. 1000 W). The software enables on-line modifications of any method parameter and the reaction process is monitored by an appropriate graphical interface. Agitation is achieved by magnetic stirring (max. 400 rpm) controlled by software.

NMR spectroscopy. All NMR spectra were recorded on a Bruker Avance spectrometer DPX 300 at 21 °C, using deuteriated propanone as solvent (10-20 mM solutions for ¹H and ³¹P NMR spectroscopy; 30-35 mM solutions for ¹³C NMR spectroscopy), referred to TMS for ¹H and ¹³C NMR spectra and 85% H₃PO₄ for ³¹P NMR spectra.

Mass spectrometry. Most samples were analysed by fast atom bombardment (FAB) spectrometry, performed on with VG-Autospec (Fisons Instruments) in a positive ion mode.

Microanalysis. Elemental carbon and hydrogen contents (wt%) were analysed using a Perkin-Elmer Series II CHNS/O 2400 CHN Elemental Analyser, and provided analysis within ±0.3 wt% error. Halide content was determined by dioxygen flask, combustion followed by halide titration with aqueous mercury(II) nitrate, using a dicarbazide indicator.

Water content. Water content was determined using a Cou-Lo Compact (version 08.05) Karl Fischer titrator, and titration solutions from Riedel-de-Haën. Ionic liquids were tested immediately after drying under high vacuum for 48 h, and the water content was below 750 ppm for all samples.

Differential scanning calorimetry (DSC). The melting points and glass transitions of the compounds were measured by DSC, using a TA Instruments Modulated DSC Q 2000 V24.4 Build 116 with refrigerated cooling system RCS 90 capable of controlling the temperature down to 220 K. Dry dinitrogen gas was purged through the DSC cell with a flow rate of *ca.* 20 cm³ min⁻¹. All

samples, due to their hygroscopic character, were prepared in a glove box, using Tzero hermetic platinum pans to avoid any contact with air and measured under the same conditions (standard heating and cooling ramp of 5 K min⁻¹ over five cycles). The melting point is taken as the onset of an endothermic peak on heating and the glass transition temperature is taken as the midpoint of a small heat flow change on heating from the amorphous glass state to a liquid state.

Thermogravimetric analysis (TGA). Decomposition temperature measurements were performed in a TA Instruments Q50 thermogravimetric analyser. The samples were measured in platinum pans, at a heating rate of 10 K min⁻¹, in a dynamic mode, under dinitrogen atmosphere. The onset of the weight loss in each thermogram was used as a measure of decomposition temperature (see discussion). The sample size was typically between 7 and 10 mg and sample preparation was similar to that described for the DSC measurements.

Density measurements. Density was measured with an Anton Paar vibrating tube densimeter, model DMA 4500, operating at atmospheric pressure and within the temperature range 273 to 363 K. The internal calibration of the instrument was confirmed by measuring densities of atmospheric air and doubly distilled water, according to the recommendations of the manufacturer. The DMA 4500 cell was embedded in a metallic block, the temperature of which was controlled by several Peltier units. This arrangement enabled a temperature stability better than ± 2 mK.

All ionic liquids used in density determinations were degassed under vacuum and moderate temperature conditions for periods longer than 48 h and were injected into the densimeter into the densimeter immediately after cooling, using non-lubricated disposable syringes.

Viscosity measurements. The viscosity was determined using an AMVn microviscometer, Anton Paar, that allows measurements from 273 to 363 K at atmospheric pressure, and over a wide viscosity range (0.3 to 20 000 mPa s). A glass capillary with 4 mm internal diameter with a 3 mm gold/tungsten ball (placed inside the capillary) with density 15 g cm⁻³ was chosen for this study and an extension of runtime up to 1000 s was required due to the high viscosity of each sample. Temperature was maintained constant to within ±0.01 K by a capillary block with in-built Peltier elements. Since these ionic liquids were extremely hygroscopic, the capillary was always charged in the glove box. The capillary/ball system was calibrated with standard viscosity oil (1 Pa s) from PRA. Viscosity measurements were performed in steps of 10 K from 293 K up to 363 K for all ionic liquid samples for which a sufficient amount (*ca.* 3.5 cm³) was available. Each sample was measured ten times in order to check reproducibility.

X-ray analysis. Crystal data for [P_{4.4.4.3}]Cl were collected using a Bruker kappaCCD diffractometer with graphite monochromated Mo-K_α radiation at *ca.* 150 K in a dinitrogen stream. Lorentz and polarisation corrections were applied. The structure was solved by direct methods and the non-hydrogen atoms were refined with anisotropic thermal parameters. Hydrogen-atom positions were added at idealised positions and a riding model with fixed thermal parameters ($U_{ij} = 1.2U_{eq}$ for the atom to which they are bonded (1.5 for methyl)), was used for subsequent refinements. The nature of the disorder in the cation conformations required that

restraints be included on all of the P-C, C-C and C-H bond lengths and bond angles to provide a chemically sensible model for the cations. The function minimised was $\Sigma[w(|F_0|^2 - |F_c|^2)]$ with reflection weights $w^{-1} = [\sigma^2 |F_0|^2 + (g_1 P)^2 + (g_2 P)]$ where $P = [\max(|F_0|^2 + 2|F_c|^2)/3]$. The SHELXTL package³⁰ was used for structure solution and refinement.

Search procedures in the Cambridge Structural Database (CSD) were carried out using ConQuest 1.13.^{31, 32} The CSD (version 5.32, updated November 2010) was examined using the following criteria: 3D coordinates determined, with no errors and no powder structures.

Single crystals of C₁₅H₃₄CIP were grown by slow crystallisation from the melt at 70 °C. Crystal data for [P₄₄₄₃]Cl. C₁₅H₃₄CIP, *T* = 150 K, monoclinic, *P*2₁/*c*, *a* = 9.9996(5), *b* = 13.4309(12), *c* = 13.3481(10) Å, β = 91.331(5)°, *V* = 1792.2(2) Å³, *Z* = 4, *D*_c = 1.041 g cm⁻³, reflections collected/unique = 7858/2329 (*R*_{int} = 0.0821), μ = 2.565 mm⁻¹, *S* = 1.065, *R*₁ = 0.0821 (*I* > 2σ), *wR*(*F*₂) = 0.2539 (all data).

Glove box. All ionic liquids studied here were handled in a dinitrogen-filled glove box (MBraun LabMaster dp; 0.1 ppm < O₂ and H₂O) immediately after drying, and stored there until used.

Materials

The phosphonium ionic liquids presented in this study were synthesised following the procedures described in the next section. Chloromethane (≥99.5%), chloroethane (≥99.7%), 1-chloropropane (98%), 1-chloropentane (99%), 1-chlorohexane (99%), 1-chloroheptane (99%), 1-chlorooctane (99%), 1-

Table 1 Masses of reagents (*m*), reaction conditions for synthesis, yield, appearance and water content of [P_{444*n*}]Cl ionic liquids, masses and volumes of P₄₄₄ reported at pure state (P₄₄₄ was supplied in a 50% solution in toluene).

<i>n</i>	P ₄₄₄		C _n H _{2n+1} Cl		$\rho/g\text{ cm}^{-3}$ *	<i>T</i> /°C	<i>t</i> /h	Yield/%	Colour	State ^c	H ₂ O/ppm
	<i>V</i> /cm ³	<i>m</i> /g	<i>V</i> /cm ³	<i>m</i> /g							
2	48.80	40	32.70	30	0.918	140 ^a	4.5 ^a	88	Pale yellow	s	648
3	7.32	6	2.90	2.60	0.892	180 ^b	0.5 ^b	93	Pale yellow	s	622
4	-	-	-	-	-	-	-	-	Pale yellow	s	593
5	14.60	12	7.90	6.95	0.882	145	12	90	White	s	543
6	14.60	12	8.94	7.90	0.879	145	12	92	Colourless	l	450
7	14.60	12	9.96	8.80	0.881	145	14	94	White	s	531
8	14.60	12	11.04	9.70	0.875	145	14	92	Colourless	l	447
10	14.60	12	13.23	11.50	0.868	145	16	90	Colourless	l	492
12	14.60	12	15.35	13.30	0.867	145	16	93	Colourless	l	417
14	-	-	-	-	-	-	-	-	Pale yellow	s	373

^a Synthesis performed in autoclave; ^b Synthesis performed in microwave reactor; ^c s – solid; l – liquid; *density data provided by supplier.

Synthesis of [P₄₄₄₁]Cl. A stoichiometric amount of tributylmethylphosphonium methylcarbonate (75% solution in methanol, 0.103 mol) and concentrated aqueous hydrochloric acid (0.103 mol) were weighed in separate beakers, 500 cm³ and 50 cm³, respectively. The acid was slowly added into the ionic liquid, while constantly stirring at room temperature, which helped to reduce the amount of foam generated by the evolved CO₂ gas. When CO₂ evolution ceased, the reaction was considered as complete. The reaction mixture was transferred to a round-bottomed flask (100 cm³) and further dried at 65 °C *in vacuo* for 48 h. The reaction yield was 91.5% (23.8g); the product was a colourless viscous liquid.

Synthesis of [P₄₄₄₂]Cl. The autoclave glass reactor was charged with a tributylphosphine solution in toluene (see Table 1) under a dinitrogen atmosphere, placed into the stainless steel cage holder, and quickly attached to the cover plate. A gas cylinder

chlorononane (98%), 1-chlorodecane (98%), 1-chlorododecane (98%), were all purchased from Aldrich and used as received. Tributylphosphine (50% solution in toluene, $\rho = 0.82\text{ g cm}^{-3}$), Tributyl(methyl)phosphonium methylcarbonate (75% solution in methanol), [P₄₄₄₄]Cl and [P₄₄₄₁₄]Cl were supplied by Cytec Industries, Canada.

Synthetic procedures

Four different techniques were used to synthesise alkyltributylphosphonium chlorides. [P₄₄₄₁]Cl was synthesised by a neutralisation reaction of methanolic tributylmethylphosphonium methylcarbonate with concentrated aqueous hydrochloric acid. An autoclave was used for the reaction with chloroethane. Synthesis of [P₄₄₄₃]Cl which involved chloropropane was performed in a microwave reactor, and the remaining compounds requiring low-volatility chloroalkanes were synthesised *via* thermal reflux. Glassware used was charged with starting materials in an M-size Atmosbag® (purchased from Sigma Aldrich) which was filled with dinitrogen, in order to provide an inert atmosphere and avoid potential oxidation of the alkylphosphines to alkylphosphine oxides. Details of individual syntheses are given in Table 1, elemental analytical results are reported in Table 2. In all cases reactions were characterised by ³¹P, ¹³C, ¹H NMR (Tables 3 and 4; Figure 2) and by mass spectrometry (Table 2) prior to purification.

containing chloroalkane (chloromethane or chloroethane) was connected to the autoclave through the gas inlet tube. The gas and inlet valves were opened after the set up was proved to be tight and the gas flow was maintained at approximately 2 cm³ min⁻¹. The reaction mixture was slowly heated over several hours in an oil bath until the temperature inside of the reactor reached 145 °C. The pressure in the reactor was built up by an excess of chloroalkane and did not exceed 4 bar with increasing temperature. The gas inlet was closed after approximately one hour and the reaction was maintained at 145 °C for two hours in the case of chloromethane, and for an additional two hours in case of chloroethane. After the reaction was cooled down to room temperature, the pressure was released and the sample for ³¹P NMR spectroscopy was taken to check for the disappearance of phosphine signal. The content of the reactor was then transferred to a round-bottomed flask (100 cm³) and dried at 65 °C *in vacuo* for 48 hours.

Synthesis of [P₄₄₄₃]Cl. A tributylphosphine solution in toluene (1 mol eq.) was mixed with 1-chloropropane (1.1 mol eq.) in a glass microwave vial (20 cm³) with an inserted stirring bar, under dinitrogen. The reaction mixture was transferred under a Suba-Seal[®] septum (purchased from Sigma Aldrich), avoiding any contact with air while transferring from the Atmosbag[®] to the shield vessel. The septum was removed and discarded immediately before the TMS screw cap was attached to the shield vessel. They were then placed inside the microwave cavity. The reaction conditions were: Ramp (5 min) to 180 °C; hold at 180 °C (30 min), medium stirring, 500 W (50% max) power; pressure reached 3 bar. After the reaction was complete and cooled, the product was transferred to a round-bottomed flask (25 cm³) and

dried at 65 °C *in vacuo* for 48 h.

Synthesis of [P_{444n}]Cl, n = 5-8, 10 or 12. Tributylphosphine (1 mol eq.) and 1-chloroalkane (1.1 mol eq.) were placed in a round-bottomed flask (50 cm³). The mixture was stirred and heated under a reflux condenser connected to a Schlenk line flushed with dinitrogen. The reaction mixture was heated in a silicone oil bath up to 145 °C and maintained for 12 to 16 h, depending on the alkyl chain length of chloroalkane used. In order to remove the excess of starting 1-chloroalkane (C_nH_{2n+1}Cl), where n ≥ 5, the purification step involved dissolution of product in water, followed by extraction with ethyl ethanoate. All products were then dried at 65 °C *in vacuo* for 48 h.

Table 2 Elemental analysis of [P_{444n}]Cl and identification of cations by FAB, where peaks were assigned as [M⁺].

n	C / %		H / %		Cl / %		P / %		m/z	
	calc.	exp.	calc.	exp.	calc.	exp.	calc.	exp.	calc.	exp.
1	61.76	61.83	11.96	11.74	14.02	13.86	12.25	12.57	217.2	216.9
2	63.02	63.35	12.09	12.12	13.29	13.61	10.50	10.92	231.2	230.9
3	64.15	63.48	12.20	11.80	12.62	13.06	11.03	11.66	245.2	245.2
4	65.16	64.55	12.31	12.64	12.02	12.21	10.50	10.60	259.3	259.2
5	66.10	65.75	12.40	12.52	11.48	11.54	10.03	10.19	273.3	273.3
6	66.95	66.65	12.48	12.64	10.98	10.96	9.59	9.75	287.3	287.2
7	67.72	66.83	12.56	12.90	10.52	11.25	9.19	9.02	301.3	301.2
8	68.44	68.23	12.64	12.13	10.10	10.32	8.82	9.32	315.3	315.3
10	69.71	68.90	12.76	12.75	9.35	9.68	8.17	8.68	343.4	343.4
12	70.81	70.89	12.87	12.23	8.71	8.77	7.61	8.11	371.4	371.6
14	71.76	71.54	12.97	12.90	8.15	8.05	7.12	7.51	399.4	399.4

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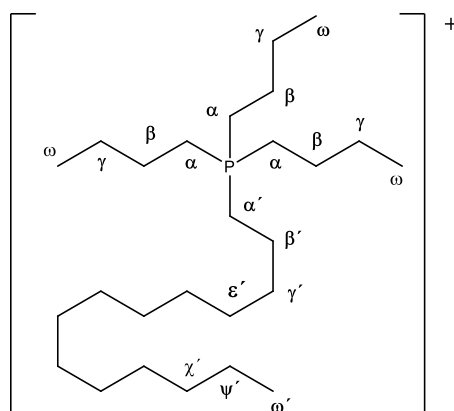


Fig. 2 The numbering convention used for structures of the cation described in **Tables 3** and **4**. The fourth (variable) alkyl substitution on the phosphorus is labelled as a secondary chain (α' , β' , γ' , etc.) where terminal position takes priority, apart from positions α' and β' .

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Table 3 For [P_{444n}]Cl, ¹H NMR shifts (δ/ppm, *d*₆-propanone, 300 MHz, 21 °C), with coupling constant (²*J*_{PH}/Hz) and ³¹P{¹H} NMR shifts (δ/ppm, *d*₆-propanone, 120 MHz, 21°C), peaks were assigned according to **Figure 2**.

<i>n</i>	¹ H ^{α,α'}		¹ H ^{β,β'}		¹ H ^{γ,γ'-ψ'}		¹ H ^{ε,ε'}		³¹ P
	δ/ppm	intensity	δ/ppm	intensity	δ/ppm	intensity	δ/ppm	intensity	
1	2.50-2.60 (m), 2.08 (d, ² <i>J</i> = 14.4)	6H,3H ^(α)	1.47-1.60 (m)	6H	1.31-1.43(m)	6H	0.81-0.86 (m)	9H	33.76
2	2.49-2.63 (m)	8H	1.48-1.60 (m), 1.11-1.22 (m)	6H,3H ^(β)	1.31-1.43 (m)	6H	0.81-0.86 (m)	9H	36.35
3	2.44-2.57 (m)	8H	1.48-1.65 (m)	8H	0.94-0.99 (m), 1.31-1.43 (m)	3H ^(γ) ,6H	0.81-0.86(m)	9H	34.19
4	2.45-2.55 (m)	8H	1.48-1.61 (m)	8H	1.31-1.43 (m)	8H	0.81-0.85 (m)	12H	34.61
5	2.47-2.58 (m)	8H	1.48-1.61 (m)	8H	1.20-1.43 (m)	10H	0.75-0.86(m)	12H	34.57
6	2.50-2.60 (m)	8H	1.48-1.61 (m)	8H	1.20-1.25 (m), 1.32-1.44 (m)	4H,8H	0.74-0.79 (m), 0.81-0.86 (m)	3H,9H	34.53
7	2.50-2.60 (m)	8H	1.48-1.61 (m)	8H	1.15-1.27 (m), 1.32-1.44 (m)	6H,8H	0.74-0.78 (m), 0.81-0.86 (m)	3H,9H	34.51
8	2.39-2.50 (m)	8H	1.38-1.51 (m)	8H	1.07-1.15 (m), 1.21-1.34 (m)	8H,8H	0.63-0.67 (m), 0.71-0.76 (m)	3H,9H	34.51
10	2.49-2.60 (m)	8H	1.48-1.61 (m)	8H	1.16-1.27 (m), 1.32-1.44 (m)	12H,8H	0.73-0.77 (m), 0.81-0.86 (m)	3H,9H	34.48
12	2.49-2.60 (m)	8H	1.48-1.63 (m)	8H	1.16-1.26 (m), 1.48-1.63 (m)	16H,8H	0.73-0.77 (m), 0.81-0.86 (m)	3H,9H	34.46
14	2.46-2.56 (m)	8H	1.48-1.61 (m)	8H	1.16-1.27 (m), 1.32-1.44 (m)	20H,8H	0.73-0.77 (m), 0.81-0.86 (m)	3H,9H	34.52

Table 4 For [P_{444n}]Cl, ¹³C{¹H} NMR shifts (δ/ppm, *d*₆-propanone, 75 MHz, 21 °C), with coupling constants (¹*J*_{PC}/Hz), peaks were assigned according to **Figure 2**

<i>n</i>	¹³ C ^{α,α'}	¹³ C ^{β,β'}	¹³ C ^{γ,γ'}	¹³ C ^{ε'-ψ'}	¹³ C ^{ε,ε'}
	δ/ppm	δ/ppm	δ/ppm	δ/ppm	δ/ppm
1	4.99(d, <i>J</i> _{PC} = 51.75 Hz), 20.99(d, <i>J</i> _{PC} = 49.13 Hz)	24.63	24.97, 25.19		14.30
2	13.60(d, <i>J</i> _{PC} = 48.90 Hz), 19.34(d, <i>J</i> _{PC} = 47.55 Hz)	24.71, 6.70	25.03, 25.24		14.23
3	19.71(d, <i>J</i> _{PC} = 47.55 Hz), 21.90(d, <i>J</i> _{PC} = 47.33 Hz)	24.73, 16.66	25.03, 25.23		14.23, 16.35
4*	19.70(d, <i>J</i>_{PC} = 47.55 Hz)	24.74	25.03, 25.23		14.23
5	19.67(d, <i>J</i> _{PC} = 47.55 Hz), 19.81(d, <i>J</i> _{PC} = 47.40 Hz)	23.10, 24.78	25.04, 25.25, 34.08, 34.28	22.43	14.26, 14.61
6	19.72(d, <i>J</i> _{PC} = 47.46 Hz), 19.91(d, <i>J</i> _{PC} = 47.33 Hz)	23.10, 24.79	25.04, 25.25	22.43, 34.08, 34.29	14.27, 14.62
7	19.74(d, <i>J</i> _{PC} = 47.48 Hz), 19.99(d, <i>J</i> _{PC} = 47.25 Hz)	23.57, 24.82	25.06, 25.27	22.80, 31.67, 31.87, 31.29	14.31, 14.81
8	19.74(d, <i>J</i> _{PC} = 47.46 Hz), 19.97(d, <i>J</i> _{PC} = 47.25 Hz)	23.76, 24.81	25.06, 25.27	22.76, 30.29, 31.99, 32.20, 32.99	14.29, 14.84
10	19.75(d, <i>J</i> _{PC} = 47.48 Hz), 19.98(d, <i>J</i> _{PC} = 47.25 Hz)	23.77, 24.83	25.06, 25.72	22.81, 30.10, 30.30, 32.00, 32.21, 33.00	14.32, 14.86
12	19.77(d, <i>J</i> _{PC} = 47.40 Hz), 20.00(d, <i>J</i> _{PC} = 47.10 Hz)	23.78, 24.85	25.08, 25.29	22.85, 30.23, 30.50, 30.75, 32.03, 32.23, 33.00	14.36, 14.89
14	19.73(d, <i>J</i> _{PC} = 47.55 Hz), 19.95(d, <i>J</i> _{PC} = 47.25 Hz)	23.76, 24.79	25.05, 25.26	22.78, 30.11, 30.36, 30.62, 30.80, 31.99, 32.19, 33.08	14.27, 14.83

*Due to the symmetry of the compound, all the peak assignments are equal.

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Results and discussion

Synthetic procedures

The salts, $[P_{444n}]Cl$, where $n = 2, 3, 5, 6, 7, 8, 10$ and 12 , were prepared by nucleophilic (S_N2) addition of tributylphosphine to the appropriate linear chloroalkane. As the tributylphosphine was supplied in 50% (v/v) toluene solution (in order to reduce flammability), the solvent had to be removed when the reaction was complete. The toluene retarded the reaction rate, such that it became close to that of the neat reactions of trioctylphosphine.¹¹

A different, and more convenient, approach was adopted for the synthesis of $[P_{4441}]Cl$ due to the availability of $[P_{4441}][CH_3OCO_2]$ (*ex* Cytec, Canada), which was prepared by the reaction between tributylphosphine and dimethylcarbonate. This can also be prepared easily in the laboratory by a procedure analogous to one described elsewhere.³³ Once formed, this is a versatile reagent, reacting with any Brønsted acid, HY, to quantitatively yield $[P_{4441}]Y$ (see **Figure 3**). In the specific case of $[P_{4441}]Cl$, the reagent was treated with concentrated aqueous hydrochloric acid.

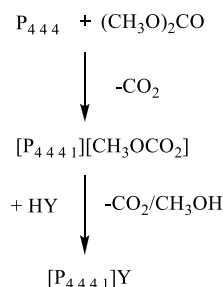


Fig. 3 A generic synthesis of tributylmethylphosphonium salts

The average yield for all the compounds reported in **Table 2** was 94 %, with the final purification being the yield reducing step, which was performed to remove the excess of chloroalkane from the product, and consisted in the dissolution of the hydrophilic ionic liquid in water followed by washing with ethyl ethanoate. Water and ethyl ethanoate were then removed on a rotary evaporator and the products were further dried at 65 °C, *in vacuo*, for 48 h. All liquid products were colourless, while solids were typically white to pale yellow waxes. All the salts $[P_{444n}]Cl$ were hydrophilic and dissolved in water.

Nuclear magnetic resonance spectroscopy

³¹P NMR spectra. The ³¹P{¹H} NMR spectra of the alkyltributylphosphonium chlorides (see **Table 3**) each exhibited a singlet at $+34.49 \pm 0.30$ ppm, with the two exceptions of $[P_{4441}]Cl$ and $[P_{4442}]Cl$, where the signals were at $+33.76$ ppm and $+36.35$ ppm, respectively. A similar phenomenon was observed for $[P_{88n}]Cl$, where the methyl and ethyl salts also exhibited different chemical shifts from the rest of the series. The singlet was accompanied by minor satellites as a result of ¹³C-³¹P coupling ($^1J_{CP} = 47.22 \pm 0.46$ Hz) to the α -carbon.

As the tertiary tributylphosphine signal is at -31.06 ppm (most of tertiary alkylphosphines exhibit signals in the δ -30 to -60 ppm range), monitoring of the reactions was very straightforward. The purity of the starting phosphine was determined by ³¹P NMR spectroscopy as *ca.* 95%. The tributylphosphine is industrially prepared (Cytec) by free radical anti-Markovnikov addition of phosphine to 1-butene,³⁴ but there is a small fraction which proceeds in classic Markovnikov mode,³⁵ where phosphorus adds to the β -carbon, forming dibutyl(*sec*-butyl)phosphine (-17.88 ppm, **Figure 4**).

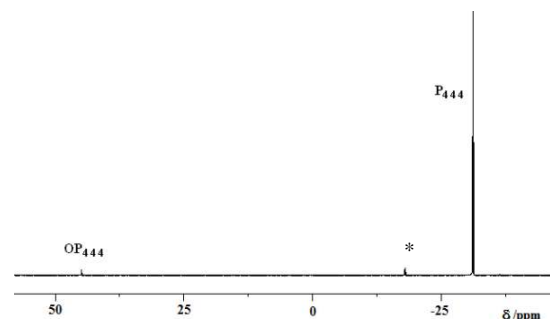


Fig. 4 The ³¹P{¹H}NMR spectrum (*d*₆-propanone, 120 MHz, 21 °C) with peak assignments for tributylphosphine, tributylphosphine oxide and dibutyl(*sec*-butyl)phosphine (marked with asterisk).

In addition, at $+44.90$ ppm, traces of tributylphosphine oxide were observed. Its presence indicates the extreme sensitivity of tributylphosphine to aerial oxidation and reinforces the need for exceptional care in handling the phosphine.

¹H and ¹³C NMR spectra. Proton and carbon-13 NMR spectra of $[P_{444n}]Cl$ are reported and assigned in **Tables 3** and **4**. A major difference in the chemical shifts of the alkyl protons can be observed with an increasing number of carbon atoms in the fourth alkyl chain on the cation. The protons of the methyl substituent in $[P_{4441}]Cl$, see **Figure 5(B)**, can be observed as a doublet due to the strong phosphorus–hydrogen coupling.

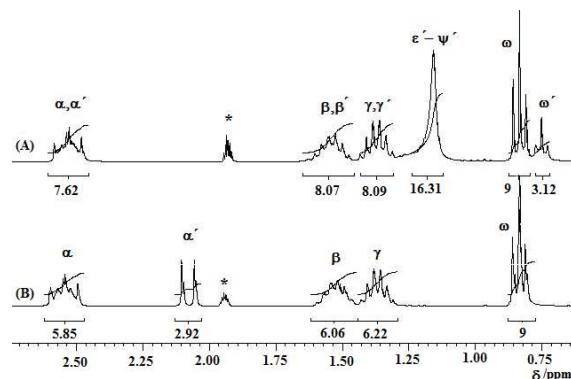


Fig. 5 The ¹H NMR spectra (*d*₆-propanone, 300 MHz, 21 °C) of (A) $[P_{44412}]Cl$ and (B) $[P_{4441}]Cl$; peaks are labelled according to **Figure 2** (individual peaks are assigned in **Table 3**), solvent is assigned as (*).

As the number of methylene groups increases, the shift of the

methyl group at the end of the fourth alkyl substitution moves towards to lower magnetic field until it reaches the multiplet belonging to the methyl groups of the three remaining butyl substitutions on the cation. The multiplet belonging to protons bonded to a terminal carbon visibly split when $n = 6$, shifting the signal of $-\text{CH}_3$ group on the fourth alkyl substituent further downfield, **Figure 5** (see **Table 3**). The sensitivity of the unique terminal methyl group is obvious in the ^{13}C NMR spectra as well, see **Table 4** and **Figure 6**. With increasing n , the signal of the terminal group approaches that of the remaining methyl groups at $n = 4$, when the compound adopts tetrahedral symmetry. As n becomes > 4 , the signal at +14.23 ppm splits into two signals and remains effectively unchanged up to $n = 14$.

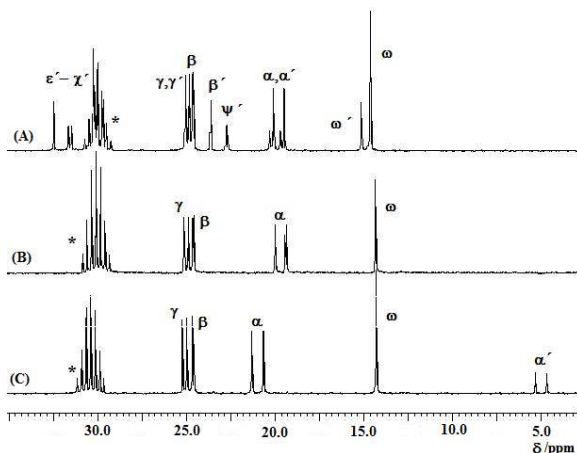


Fig. 6 The $^{13}\text{C}\{^1\text{H}\}$ NMR spectra (d_6 -propanone, 120 MHz, 21°C) of (A) $[\text{P}_{4.4.4.8}]\text{Cl}$, (B) $[\text{P}_{4.4.4.4}]\text{Cl}$ and (C) $[\text{P}_{4.4.4.1}]\text{Cl}$; peaks are labelled according to **Figure 2** (individual peaks are assigned in **Table 4**), solvent is assigned as (*).

In order to justify the assignments for the NMR spectra, show in **Tables 3** and **4**; and **Figure 2**, 2-D COSY and 2-D HSQC spectra, respectively, see **Figures 7** and **8**, are presented for $[\text{P}_{4.4.4.3}]\text{Cl}$.

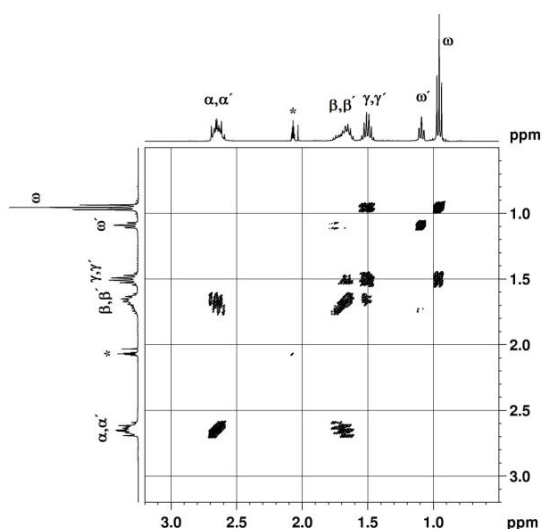


Fig. 7 The 2-D ^1H , ^1H -COSY NMR spectrum of $[\text{P}_{4.4.4.3}]\text{Cl}$ (d_6 -propanone, 400 MHz, 21°C), solvent is assigned as (*) and peaks according to **Figure 2**.

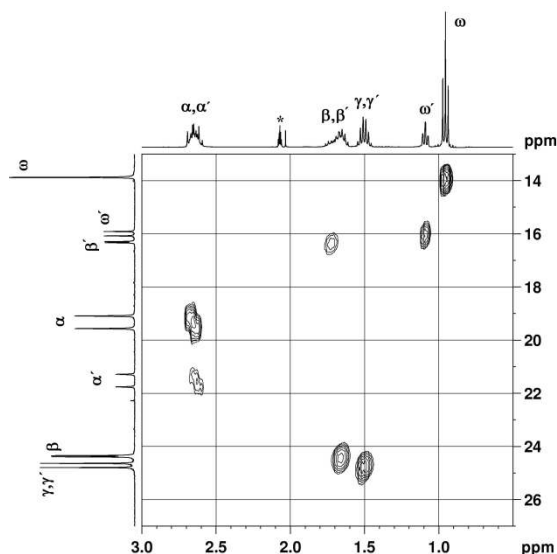


Fig. 8 The 2-D ^1H , $^{13}\text{C}\{^1\text{H}\}$ -HSQC NMR spectrum of $[\text{P}_{4.4.4.3}]\text{Cl}$ (d_6 -propanone, 400 MHz, 21°C), solvent is assigned as (*) and peaks labelled according to **Figure 2**.

These experiments serve as a justification for assignment of chemical shifts, especially in ^{13}C NMR. The chemical shifts, when assigned according to ^{13}C NMR predictor (ChemBioDraw Ultra[®]),³⁶ were in excellent agreement with results obtained from 2-D HSQC NMR experiment.

Single crystal analysis

The crystal structure of $[\text{P}_{4.4.4.3}]\text{Cl}$ has been determined by X-ray diffraction, which confirmed the absence of any solvation. This salt crystallised from its melt in a monoclinic lattice, space group $P2_1/c$, with one ion pair in the asymmetric unit (see experimental section). The cation has been refined with a positional disorder of two components (50/50 site occupancy factor ratio), as seen in **Figure 9**. For both cation components, the central phosphorus atoms have been modelled as occupying identical positions.

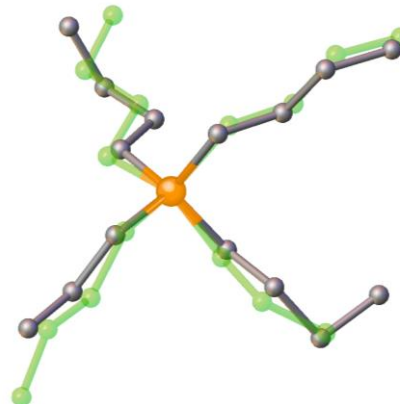


Fig. 9 A superimposed view of both disordered cation fragments, AA and AG, in the crystal structure of $[\text{P}_{4.4.4.3}]\text{Cl}$. Orange = phosphorus; the carbon atoms are coloured grey (AG) and green (AA) to distinguish the cations; all hydrogen atoms have been omitted for clarity.

When comparing both cation conformations, the propyl chain and one of the butyl chains are located in exchanged regions and these two conformations are randomly distributed throughout the

lattice. The conformations adopted within the constraints are consistent with known quaternary phosphonium salts, and the average P-C bond distance in both cation fragments of $[P_{4443}]Cl$ structure is $1.80 \pm 0.03 \text{ \AA}$, in good agreement with those previously reported, $1.79 \pm 0.02 \text{ \AA}$.³⁷ In one conformation, the alkyl chains display all-*anti* conformations, **AA**, whereas the other conformation has two butyl chains with their terminal methyl groups in a *gauche* conformation, *i.e.* the bonds between the third and fourth carbon atoms show marked torsions (dihedral angles = 64 and 83° , see **Figure 9**), **AG**.

By detailed examination of the crystal structure of $[P_{4443}]Cl$, it has been observed that each cation is surrounded by three chloride anions; see **Figure 10(A)**. Conversely, each anion is surrounded by three cations; **Figure 10(B)**. A number of short contacts have been found between the chloride anion and hydrogen atoms in the α position to the phosphorus atom (**Figure 10**). The distances of such contacts in conformer **AA** are 2.62 , 2.70 , 2.78 , 2.79 and 2.85 \AA , giving an average of $2.75(9)$, and in conformer **AG** are 2.48 , 2.73 , 2.83 , 2.84 and 2.91 \AA giving an average of $2.76(17) \text{ \AA}$: these contacts are significantly shorter than the sum of Van der Waals radii (2.95 \AA).^{38,39} This suggests that, in addition to the strong electrostatic attraction between the chloride anions and the positively charged phosphonium cations, hydrogen bonding should also be of significance in the solid state of $[P_{4443}]Cl$. Furthermore, as expected, the $P \cdots Cl$ distances were found in the 3.93 - 4.34 \AA range, well above the sum of Van der Waals radii (3.55 \AA).^{38,39}

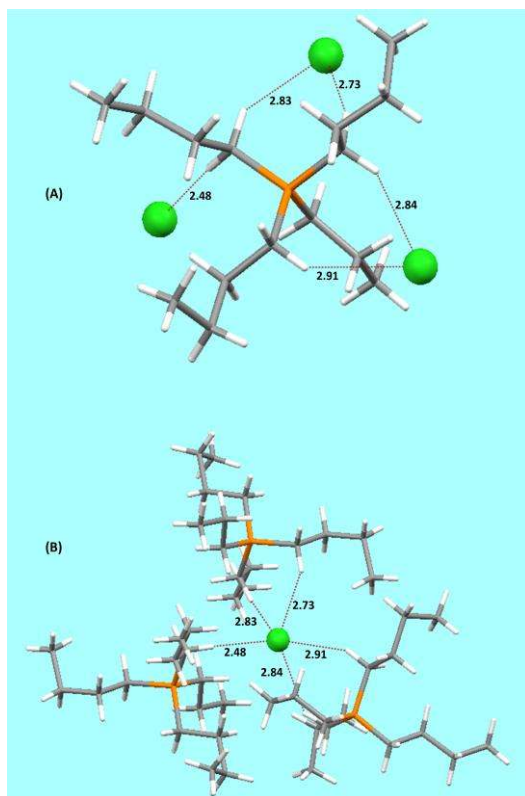


Fig. 10 (A) The anion coordination around the phosphonium cation and (B) the cation coordination around the chloride anion in the crystal structure of $[P_{4443}]Cl$. For clarity, only one of the cation components (**AG**) is shown. Orange = phosphorus, grey = carbon, white = hydrogen, and green = chlorine. $Cl \cdots H$ contacts (\AA) are represented by dashed red lines.

As seen in **Figure 11**, the chloride anions and the phosphorus atom of the cations in $[P_{4443}]Cl$ form layers parallel to the crystallographic bc plane throughout the structure. The methyl ends of the alkyl chains occupy the boundaries of the segregated interlayers. Such a mode of packing resembles those found in the lattices of $[N_{4444}]Br$ ⁴⁰ and $[N_{4444}]I$,⁴¹ both these salts crystallising in monoclinic lattices, as is the case for $[P_{4443}]Cl$. This is not surprising, given the structural similarity of these salts with the one reported here.

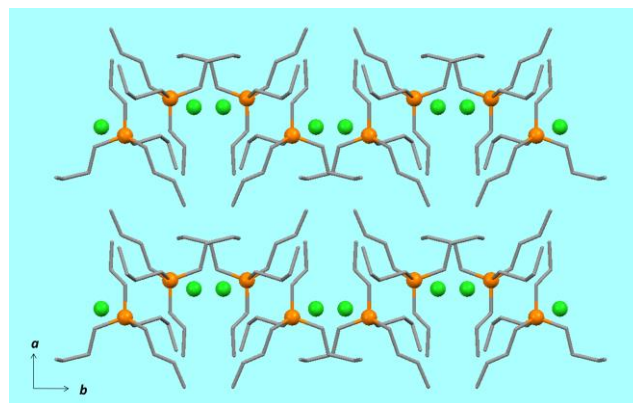


Fig. 11 A view of the crystal packing in the crystal structure of $[P_{4443}]Cl$, showing the arrangement into layers parallel to the crystallographic bc plane. The colours are as in **Figure 10**. For clarity, only one of the cation components (**AG**) is shown and hydrogen atoms have been omitted.

The crystal structure of $[P_{4443}]Cl$ reported here is, to the best of our knowledge, the first example of a structure of an anhydrous chloride salt of a phosphonium cation bearing only alkyl substituents. It should be noted that the crystal structures of tris(cyclohexyl)methylphosphonium chloride trihydrate⁴² and tetramethylphosphonium chloride monohydrate⁴³ have been reported, where the structures are dominated by strong hydrogen bonds between water and the chloride ions. Anhydrous analogues for other halides have been structurally determined, namely tetramethylphosphonium fluoride ($[P_{1111}]F$),⁴⁴ tetrakis(decyl)phosphonium bromide ($[P_{101010}]Br$),⁴⁵ and a small number of alicyclic phosphonium iodides.⁴⁶⁻⁴⁸ Furthermore, different polymorphs of anhydrous tetraethylammonium^{49, 50} and 1-methyl-1-propylpyrrolidinium^{51, 52} chlorides are also known.

A comparison between interionic $H \cdots X$ and $E \cdots X$ ($E = N$ or P , $X = \text{halide}$) distances for the salt reported here and for previously reported phosphonium and ammonium salts is shown in **Table 5**. Although the number of structures is rather limited, especially for tetraalkylphosphonium halides, some patterns become evident:

- (1) Strengthening of the hydrogen bonds is observed in the sequence $F > Cl > Br > I$, as expected.
- (2) Increasing anion volumes in the order $F < Cl < Br < I$ cause, not surprisingly, an increase in the interionic distances.
- (3) Steric encumbrance around the central phosphorus or nitrogen atoms in the cations greatly affects the cation-anion separation distances; those cations containing methyl and/or cyclised substituents are demonstrably less hindered, allowing a much closer approach of the halide anions to the central phosphorus or nitrogen atoms.
- (4) The larger size of the phosphorus atom, as compared to nitrogen, may lead to a more open environment around itself, and to a closer approach of the halide.

Table 5 Average interionic distances^a in the crystal structure of [P₄₄₄₃]Cl and in related fully alkylated phosphonium or ammonium halides^c.

	average <i>d</i> (C _α H... X) ^a	average <i>d</i> (C _β H... X) ^a	minimum <i>d</i> (E...X) ^b
Phosphonium salts (E = P)			
[P ₁₁₁₁]F ^d	2.27(9)	-	3.385(1)
[P ₄₄₄₃]Cl ^e	2.75(13)	-	3.955(2)
[P ₁₀₁₀₁₀₁₀]Br ^f	-	2.95(2)	4.858(2)
[P _{cyclo-4 cyclo-4}]I ^g	3.175	3.175	4.531(2)
[P _{cyclo-4 2 2}]I ^h	2.97	-	4.628(2)
Ammonium salts (E = N)			
[C _{3mpyrr}]Cl (cyclic) ^{i,j}	2.78(5)	2.78(2)	3.892(1) <i>3.89-3.91</i>
[N ₂₂₂₂]Cl ^k	2.81(10)	2.89(5)	4.278(1) <i>4.27-4.37</i>
[N _{cyclo-4 cyclo-4}]Br (spiro) ^l	2.853	2.98(6)	4.120(6)
[C _{nmpyrr}]Br (cyclic) ^{j,m}	2.94(8)	2.98(3)	4.235(4) <i>4.23-4.26</i>
[N ₁₁₁₁]Br ⁿ	2.976	-	4.373(1)
[N _{11nm}]Br ^o	2.90(9)	2.986	4.065(5) <i>4.06-4.20</i>
[NR ₄]Br (R>Me) ^{p,q}	2.90(9)	2.97(7)	4.369(8) <i>4.36-4.94</i>
[C _{nmpyrr}]I or [C _{1mpip}]I (cyclic) ^r	3.13(5)	-	4.296(1) <i>4.29-4.74</i>
[N ₁₁₁₁]I ^s	3.17	-	4.563(1)
[NR ₄]I (R>Me) ^{q,t}	3.13(3)	3.13(4)	4.645(2) <i>4.64-5.05</i>

^a Distances (in Å) averaged for all contacts shorter than the sum of Van der Waals (VdW) radii (refs. 53 and 54). ^b These distances represent the closest approach of the halide ion to the phosphorus or nitrogen from each set of compounds considered, *i.e.* they are not average values; for data sets of more than one structure, the range of minimum distances is given in italics.

^b Data retrieved from the Crystallographic Structural Database (CSD) are referred to by their six-letter code. ^d #UJJIAY. ^e This work. ^f #XISCUX.

^g [P_{cyclo-4 cyclo-4}]⁺ = 5-phosphoniaspiro[4.4]nonane; #EFOWUR.
^h [P_{cyclo-4 2 2}]⁺ = 1,1-diethylphospholanium; #DUBNAP. ⁱ #LIVWUJ, #LIVWUJ01. ^j [C_{nmpyrr}]⁺ = 1-alkyl-1-methylpyrrolidinium; [C_{nmpip}]⁺ = 1-alkyl-1-methylpiperidinium. ^k #HIVROT, #HIVROT01. ^l [N_{cyclo-4 cyclo-4}]⁺ = 5-azoniaspiro[4.4]nonane; #HIWIKUU. ^m #KUKSOZ, #YUHFUD, #YUHGAK. ⁿ #ZZZUQO03. ^o #BIFDAV, #CANYOG, #CEMEBR01, #OJIGOD, #OJIGOD01. ^p #CANYUM, #RABTIX, #TPRAMB, #TUDQEO, #XISDIM. ^q R = linear alkyl chain longer than methyl. ^r #QOPZOK, #REKYOW, #WAWHIM. ^s #QQQCVG01. ^t #ALEXEU, #CANZAT, #CANZEX, #CAVXEC01, #HEVMEA, #RABTET01, #XISDOS.

Average H...X distances for all contacts shorter than the sum of Van der Waals radii are listed in **Table 5** for both ammonium and phosphonium crystal structures. The overall trend suggests that some degree of hydrogen bonding is present in the structures, decreasing in the F⁻ > Cl⁻ > Br⁻ > I⁻ order. This is consistent with previous statistical studies performed on all structures in the CSD.⁵⁴ In the case of iodide, mean H...I distances were found in the 2.97-3.18 Å range, only marginally below the Van der Waals cut-off distance (3.18 Å), thus indicating very weak hydrogen bonding.

The effect of increasing anion size in the F⁻ < Cl⁻ < Br⁻ < I⁻ order is clearly reflected in the increasing E...X (E = N or P, X = halide) distances, as seen in **Table 5**. For example, by inspecting the minimum N...X distances for tetraalkylammonium halides, an increase in the shortest interionic separations (*ca.* 3.89, 4.06 and 4.30 Å for Cl⁻, Br⁻ and I⁻, respectively) is evident

with increasing anion size, as expected. As for tetraalkylphosphonium salts, a similar trend is observed, with the exception of [P₁₀₁₀₁₀₁₀]Br (*vide infra*), for which the minimum P...Br distance is unexpectedly long (*ca.* 4.86 Å).

For [P₄₄₄₃]Cl, the shortest P...Cl distance found is *ca.* 4.0 Å, 0.45 Å above the sum of the corresponding Van der Waals radii. The closest phosphorus and bromine atoms in [P₁₀₁₀₁₀₁₀]Br are *ca.* 4.86 Å apart, more than 1.2 Å longer than the sum of the Van der Waals radii (this is comparable to the N...Br contact in the [N₁₂₁₂₁₂₁₂]Br of *ca.* 4.58 Å). Furthermore, whereas contacts between chloride anions and hydrogen atoms in the α positions in the cation were found for [P₄₄₄₃]Cl, only contacts between bromide anions and hydrogen atoms in the β positions were observed for [P₁₀₁₀₁₀₁₀]Br. As the linear alkyl chains increase in length, two competing effects are realised: (i) the electrostatic attraction between the cation and the anion is reduced and (ii) the Van der Waals repulsive forces are increased. In other words, when the chain lengths are below say hexyl, the structure is best described as a simple salt dominated by Coulombic forces, and when the chain lengths above say octyl the structure has a much more covalent nature. In addition the shorter (and disordered) butyl and propyl chains in [P₄₄₄₃]Cl are more flexible, as opposed to the rigid, highly ordered (all-*anti* conformation) decyl chains in [P₁₀₁₀₁₀₁₀]Br.⁴⁵

Unfortunately, no structures (with 3D coordinates available) for tetraalkylphosphonium iodides (where the alkyl groups are linear) were found in the CSD. However, some crystal structures of alicyclic phosphonium salts were found.⁴⁶⁻⁴⁸ In these structures, the iodide anions are relatively close to the phosphorus atom (see **Table 5** for minimum P...I distances). If we examine the pair of iodide salts containing spiro and single heterocyclic substituents, the iodide can approach closer to the centre of the spiro compound (4.53 vs. 4.63 Å). It would therefore be expected that the unconstrained tetraalkylphosphonium iodide would have a separation of about 4.73 Å, much shorter than the analogous bromide at 4.86 Å. In spite of the larger ionic volume of iodide as compared to bromide, the former appears to approach much closer to the phosphorus atom than the latter.

Data for tetraalkylammonium halides further confirm the aforementioned effects of the alkyl substitution and/or the ring strain on the interionic separation. For example, the shortest N...Cl distances in [N₂₂₂₂]Cl and [C_{3mpyrr}]Cl ([C_{3mpyrr}]⁺ = 1-methyl-1-propylpyrrolidinium) are significantly different (*ca.* 4.36 and 4.28 Å), despite the fact that each cation type contains the same number of carbon atoms. This demonstrates that the presence strained rings around the central atom allows for a closer approach of the halide anion. These trends are further confirmed for bromide and iodide salts, where sufficiently large amount of crystal structures were found in the CSD. These data have been arranged into groups owing to the nature of the *N*-alkyl substituents (spiro vs. cyclic vs. non-cyclic, and for non-cyclic, according to the number of *N*-methyl substituents), as can be seen in **Table 5**. For cyclic cations, the minimum N...halide distances found reveal the same trend as discussed above for phosphonium iodides: the larger the ring strain, the closer the halide approach to the nitrogen atom. For example, whereas for the spiro compound [N_{cyclo-4 cyclo-4}]Br⁵⁵ (see **Table 5** for cation name) the bromide anion is as close as 4.120 Å, the range of minimum

N...Br distances for systems containing only one cycle, namely [C_nmpyrr]Br salts,^{56, 57} is 4.23-4.26 Å. In the case of non-cyclic tetraalkylammonium bromides, the bromide anion is found at minimum distances between 4.06 and 4.37 Å from the nitrogen atom for cations bearing methyl substituents, whereas this separation is increased to a 4.36-4.94 Å range for cations with linear alkyl chains longer than methyl. In the case of [N₁₁₁₁]Br, the shortest N...Br distance is found at 4.373 Å,⁵⁸ a separation noticeably longer than for other N-methyl containing salts (4.06-4.10 Å range, see **Table 5**). The more sterically compact and symmetrical tetramethylammonium cation prevents the halide ion approaching too close to the nitrogen, compared to the situation with longer more flexible chains.

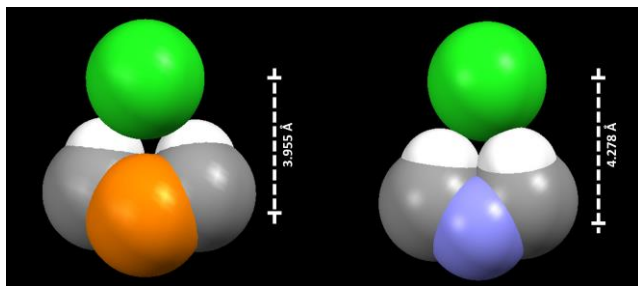


Fig. 12 Space-filling (1 Van der Waals radii) views of E(CH₂)₄-X moieties (E = P or N) in the crystal structures of [P₄₄₄₃]Cl (right) and [N₂₂₂₂]Cl (left). The rest of the cation molecules have been omitted for clarity. Orange = phosphorus, grey = carbon, white = hydrogen, blue = nitrogen and green = chlorine. E...Cl distances are schematically represented by dashed white lines.

Finally, by comparing the E...Cl distances in the structures of [P₄₄₄₃]Cl, [N₂₂₂₂]Cl and [C₃mpyrr]Cl (see **Table 5**), a further, more subtle, pattern is observed. In the structures of [N₂₂₂₂]Cl (two polymorphs), the shortest N...Cl distances are *ca.* 4.28 and 4.36 Å. In the case of [C₃mpyrr]Cl (again two polymorphs), the shortest N...Cl separations are *ca.* 3.89 and 3.91 Å [*cf.* P...Cl at 3.955 Å in [P₄₄₄₃]Cl], thus presenting a lower steric hindrance than either [N₂₂₂₂]⁺ or [P₄₄₄₃]⁺. This fact might initially look counterintuitive, given the smaller size of nitrogen as compared to phosphorus. However, a close examination of the environments around nitrogen and phosphorus atoms in [N₂₂₂₂]Cl and [P₄₄₄₃]Cl revealed that the smaller size of nitrogen is responsible for a relatively compact N(CH₂)₄ central moiety in the former, as illustrated in **Figure 12**, where the atoms occupy their full Van der Waals volume. The H_α...H_{α'} separation in the ammonium cation (for those hydrogen atoms shown in **Figure 12**) is 2.29 Å, effectively blocking the approach of the chloride anion to the nitrogen atom despite the presence of hydrogen bonding. In contrast, the H_α...H_{α'} separation in the phosphonium cation (also for those hydrogen atoms shown in **Figure 12**) is 2.69 Å, allowing a much closer approach of the chloride ion despite (or because of) the longer carbon-phosphorus bond. In other words, the more open structure around the phosphorus allows a deeper penetration of the halide ion through the protective “shrubby” of the alkyl groups than is permitted by the closed structure around the nitrogen atom. These crystallographic observations were anticipated by an elegant computational study by Hunt and co-workers.⁵⁹

50 DSC and TGA measurements

The melting (T_m) and glass transition temperatures (T_g) for [P_{444n}]Cl were measured by differential scanning calorimetry, see **Table 6**. The trends of T_m and T_g as a function of n are illustrated in **Figure 13**. All compounds (with the one exception, [P₄₄₄₁₀]Cl), exhibited distinct melting points with four of them very close to, or just below, room temperature, see **Figures 14(A)** to **14(C)**. Ionic liquids, with $n = 1$ or 5-14, also displayed glass transitions. Compounds with $n = 1, 5$ or 7 exhibited crystallisation above their glass transitions, see **Fig. 14(A)**. [P₄₄₄₁₀]Cl, as the only exception, formed a glass at -50.4 °C with no apparent freezing or melting transitions see **Fig. 14(D)**.

Table 6 Decomposition temperatures, T_d for [P_{88n}]Cl were determined by dynamic TGA. T_m and T_g obtained by DSC are reported along with the enthalpies of fusion, ΔH_{fus} , and enthalpies of glass transition, ΔH_{glass} .

n	$T_m/^\circ\text{C}$	$\Delta H_{fus}/\text{J g}^{-1}$	$T_g/^\circ\text{C}$	$\Delta H_{glass}/\text{J g}^{-1}$	$T_d/^\circ\text{C}$
1	20.4	9.66	-52.7	1.07	311
2	70.1	32.16	-	-	332
3	78.0	58.75	-	-	335
4	70.6	63.10	-	-	352
5	47.6	21.74	-40.4	0.01	361
6	24.2	2.78	-49.9	0.85	357
7	53.3	75.43	-41.2	1.37	358
8	19.4	0.23	-45.5	1.05	358
10	-	-	-50.4	1.31	338
12	22.8	49.40	-62.2	1.04	352
14	38.9	65.60	-58.8	0.64	345

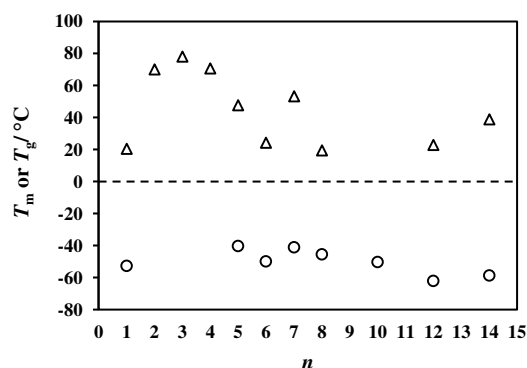


Fig. 13 Melting (Δ) and glass transition (\circ) temperatures, T_m and T_g , respectively, as a function of the number of carbons, n , in the fourth alkyl substituent on the phosphonium cation.

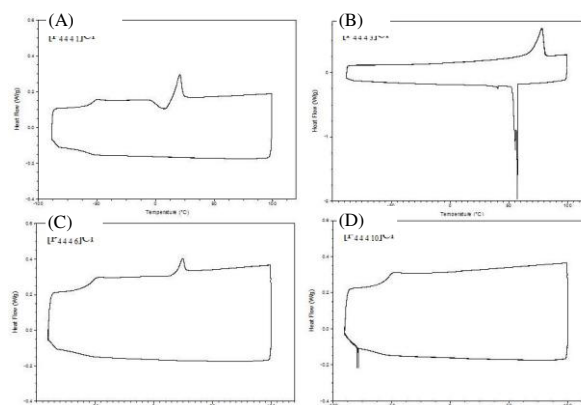


Fig. 14 DSC scans of selected compounds representing typical DSC curves for whole series of $[P_{4.44n}]Cl$ ionic liquids, (A) $[P_{4.441}]Cl$, (B) $[P_{4.443}]Cl$, (C) $[P_{4.446}]Cl$ and (D) $[P_{4.4410}]Cl$.

The thermal decomposition temperatures are listed in **Table 6**, for $[P_{4.44n}]Cl$. While the decomposition temperatures of phosphonium ionic liquids vary with the counter ion,⁷ the alkyltributylphosphonium chlorides under investigation here appear to show dynamic thermal stability around 342 ± 17 °C (four TGA scans are shown in **Figure 15**, as examples).

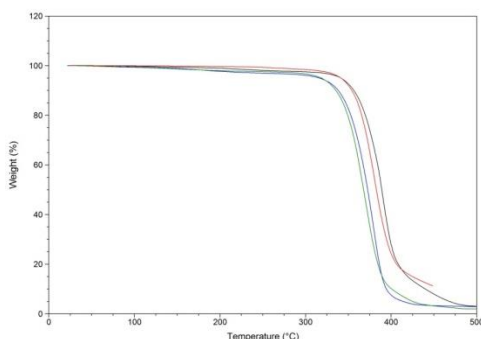


Fig. 15 A typical TGA curves for $[P_{4.44n}]Cl$, examples of $[P_{4.442}]Cl$ (black line), $[P_{4.444}]Cl$ (blue line), $[P_{4.447}]Cl$ (red line) and $[P_{4.4414}]Cl$ (green line).

Figure 16 illustrates the trend of onset decomposition temperature as a function of n , which on average is increasing with increasing alkyl chain up to $n = 5$.

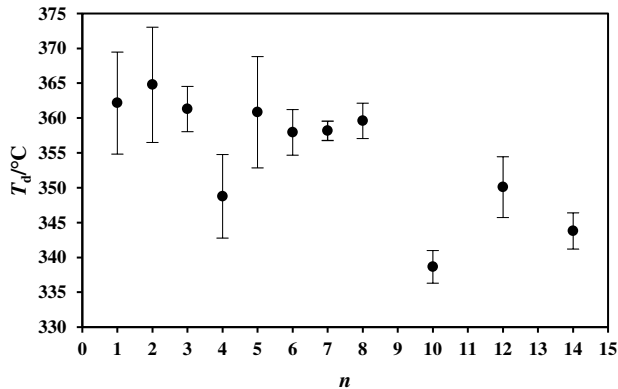


Fig. 16 Onset decomposition temperatures of $[P_{4.44n}]Cl$ ionic liquids as a function of n .

This reflects, as seen in the crystallographic section the greater resistance to the chloride ion approaching the α carbon atom. As the decomposition mechanism involves an S_N2 reaction (reversing its formation from the free phosphine and chloroalkane), the difficulty of approach of the chloride to the α carbon results in a more thermally stable material. For chain lengths where $n > 6$ the distance of closest approach reaches a plateau, as does the thermal stability. As $n > 8$ the cation becomes slightly less stable with the transparent anomaly at $n = 10$, for which we have no explanation.

Tsunashima *et al.*^{3, 60} studied two cations found in this work, ($[P_{4.441}]^+$ and $[P_{4.448}]^+$), paired with dicyanamide and bistriflamide. These ionic liquids exhibited lower melting points

and approximately 40 °C higher decomposition temperatures than their chloride analogues, as the bulky $[N(CN)_2]^-$ and $[NTf_2]^-$ anions are unable to penetrate so close to the inner carbon atoms, and are also less nucleophilic than the chloride ion.

Density and molar volume

Density measurements upon $[P_{4.44n}]Cl$, where $n = 1, 6, 8, 10$ and 12 (selected as the ones liquid at room temperature) were carried out over a broad range of temperatures ($293.15 < T/K < 363.15$) at atmospheric pressure, and data are reported along with the corresponding molar volume, V_m , **Table 7**. The repeatability of the density measurements was better than $5 \times 10^{-5} \text{ g cm}^{-3}$.

These densities are in the range of 0.87 to 0.96 g cm^{-3} over a temperature range of 293.15 to 363.15 K. All the studied ionic liquids have densities less than water and all decrease with increasing n , see **Figure 17**, and with increasing T , see **Figure 18**.

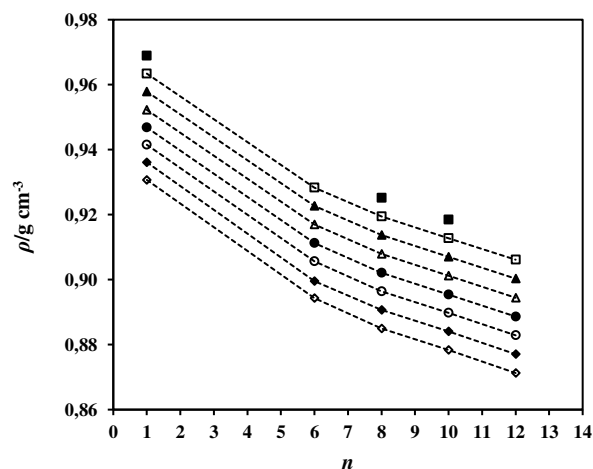


Fig.17 Experimental densities of $[P_{4.44n}]Cl$ ionic liquids at 293 K (■), 303 K (□), 313 K (▲), 323 K (△), 333 K (●), 343 K (○), 353 K (◆) and 363 K (◇) as a function n .

Experimental density data were fitted into the extended Ye and Shreeve equation⁶¹ proposed by Gardas and Coutinho,⁶² which is applicable over a wide range of temperatures 273 - 393 K, and pressures, 0.10 - 100 MPa, according to **Eq. (1)**:

$$\rho_{IL}(T, P) = M/N_A V(a + bT + cP) \quad (1)$$

where M is the formula mass in g mol^{-1} , N_A is Avogadro's number, $V^{\text{molecular}}$ is molecular volume in units of $1 \times 10^{-27} \text{ cm}^3$, T is the temperature in K, and P is the pressure in MPa.

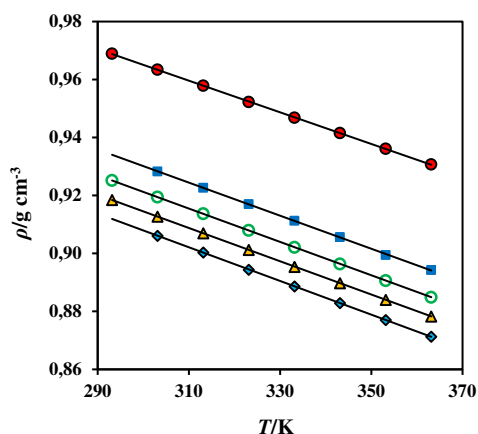


Fig. 18 Experimental density as a function of temperature for the $[P_{4.4.4.n}]Cl$ ionic liquids series where $n = 1$ (●), 6 (■), 8 (○), 10 (▲) and 12 (◆). The lines correspond to the linear fit of experimental results obtained in this work ($0.99993 \leq R^2 \leq 1$).

The values of the coefficients a , b and c were estimated by Gardas and Coutinho⁶² as 0.80050 ± 0.00023 , $(6.6520 \pm 0.0069) \times 10^{-4} \text{ K}^{-1}$, and $(-5.919 \pm 0.024) \times 10^{-6} \text{ MPa}^{-1}$, respectively, by fitting Eq.(1) to the previously published experimental pressure-volume-temperature (PVT) data of ionic liquids (imidazolium, pyridinium, pyrrolidinium and phosphonium based).^{63, 64}

The densities of alkyltributylphosphonium ionic liquids were estimated according to Eq. (1) with the molecular volumes of ions and groups taken from previously reported values^{10, 11, 62, 65} where available.

The calculated density ρ_{calc} of the studied ionic liquids displays a good agreement with the corresponding experimental density ρ_{exp} , where (see Figure 19):

$$\rho_{\text{calc}} = (0.9982 \pm 0.0015) \rho_{\text{exp}} \quad (R^2 = 0.9964 \text{ at a } 95 \% \text{ level of confidence})$$

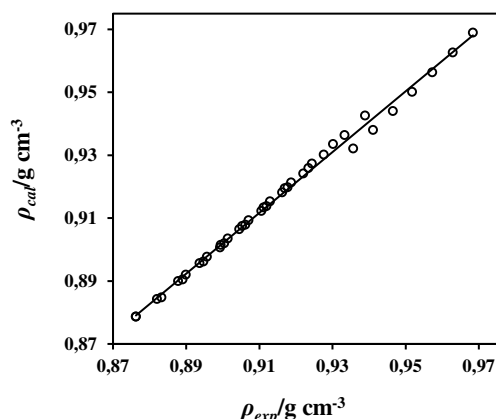


Fig. 19 Linear relationship between experimental and calculated density of herein studied ionic liquids at temperatures $T = 293.15$ - 363.15 K, using the Eq.1

Table 7 Comparisons between experimental density $\rho/g \text{ cm}^{-3}$ and molar volume data for $[P_{4.4.4.n}]Cl$ ionic liquids measured at temperatures $T = 293.15$ - 363.15 K. Calculated density values were obtained from Eq.(1). ΔV_m values represent the difference (cal - exp) between V_m values and molar volume data calculated using the fitting Eq.(6).

$T/$ K	$\rho_{\text{exp}}/$ g cm^{-3}	$\rho_{\text{calc}}/$ g cm^{-3}	$\text{Dev}^a /$ %	$(V_m)_{\text{exp}}/$ $\text{cm}^3 \text{ mol}^{-1}$	$(V_m)_{\text{calc}}/$ $\text{cm}^3 \text{ mol}^{-1}$	$\Delta V_m^b/$ $\text{cm}^3 \text{ mol}^{-1}$
[P_{4.4.4.1}]Cl						
303.15	0.96340	0.9640	0.16	262.14	256.15	-3.47
313.15	0.95786	0.9576	0.06	263.66	258.68	-2.46
323.15	0.95225	0.9514	-0.02	265.21	262.20	-1.49
333.15	0.96820	0.9452	-0.09	266.73	263.73	-0.48
343.15	0.94151	0.9391	-0.17	268.24	266.26	0.54
353.15	0.93611	0.9330	-0.26	269.79	268.78	1.52
363.15	0.93069	0.9271	-0.33	271.36	271.31	2.48
[P_{4.4.4.6}]Cl						
293.15	-	-	-	-	-	-
303.15	0.92830	0.9313	0.33	347.60	344.86	-2.75
313.15	0.92264	0.9252	0.28	349.74	347.79	-1.94
323.15	0.91699	0.9191	0.23	351.89	350.73	-1.17
333.15	0.91125	0.9132	0.21	354.11	353.66	-0.45
343.15	0.90560	0.9072	0.18	356.32	356.59	0.28
353.15	0.89950	0.9014	0.21	358.73	359.53	0.79
363.15	0.89428	0.8957	0.15	360.83	362.46	1.63
[P_{4.4.4.8}]Cl						
293.15	0.92516	0.9285	0.36	379.11	376.23	-2.88
303.15	0.91942	0.9224	0.32	381.48	379.33	-2.15
313.15	0.91368	0.9163	0.28	383.88	382.43	-1.45
323.15	0.90792	0.9103	0.26	386.31	385.52	-0.79
333.15	0.90211	0.9044	0.25	388.80	388.62	-0.18
343.15	0.89639	0.8985	0.24	391.28	391.72	0.44
353.15	0.89067	0.8927	0.23	393.79	394.81	1.02
363.15	0.88493	0.8870	0.24	396.35	397.91	1.56
[P_{4.4.4.10}]Cl						
293.15	0.91844	0.9209	0.27	412.43	410.54	-1.89
303.15	0.91272	0.9148	0.23	415.01	413.80	-1.21
313.15	0.90698	0.9088	0.20	417.64	417.06	-0.58
323.15	0.90122	0.9028	0.18	420.31	420.32	0.01
333.15	0.89543	0.8970	0.17	423.03	423.58	0.56
343.15	0.88974	0.8912	0.16	425.73	426.84	1.11
353.15	0.88402	0.8854	0.16	428.49	430.10	1.62
363.15	0.87831	0.8798	0.17	431.27	433.36	2.09
[P_{4.4.4.12}]Cl						
293.15	-	-	-	-	-	-
303.15	0.90614	0.9084	0.28	448.98	448.28	-0.71
313.15	0.90030	0.9024	0.24	451.89	451.70	-0.20
323.15	0.89443	0.8965	0.23	454.86	455.12	0.26
333.15	0.88864	0.8907	0.23	457.82	458.54	0.72
343.15	0.88286	0.8849	0.23	460.82	461.97	1.15
353.15	0.87706	0.8792	0.25	463.87	465.39	1.52
363.15	0.87126	0.8736	0.27	466.96	468.81	1.86

^a Dev (%) = $[(\rho_{\text{calc}} - \rho_{\text{exp}}) / \rho_{\text{exp}}] \times 100$; ^b $\Delta V_m = (V_m)_{\text{calc}} - (V_m)_{\text{exp}}$

For the density prediction from molar volume, using Eq.(1), an average absolute deviation (AAD) of 0.23 % with the maximum deviation (MD) 0.39 % was observed for 37 density data points (N_p), studied of 5 ionic liquids, where AAD is defined as, Eq. (2):

$$\text{AAD}(\%) = \sum_{i=1}^{N_p} \left| \frac{\rho_{\text{calc}} - \rho_{\text{exp}}}{\rho_{\text{exp}}} \right| / N_p \quad (2)$$

The volumetric behaviour of phosphonium ionic liquids can be expressed in terms of their molar volume, $V_m = M/\rho$, where an expected increase of V_m with increasing temperature is observed, see Figure 20. V_m as a function of n exhibits a quasi-linear trend for each represented temperature, Figure 21.

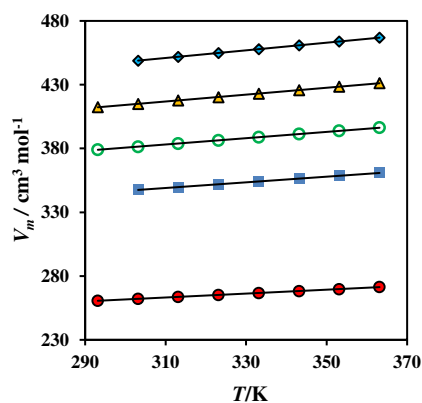


Fig. 20 Molar volumes as a function of temperature for the $[P_{4.4n}]Cl$ ionic liquids series where $n = 1$ (●), 6 (■), 8 (○), 10 (▲) and 12 (◆). The lines correspond to the linear fit of experimental results obtained in this work ($0.999735 \leq R^2 \leq 0.99997$).

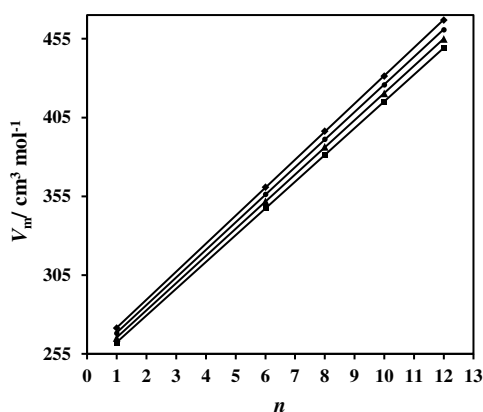


Fig. 21 Experimental molar volume of $[P_{4.4n}]Cl$ ionic liquids as a function of n for four temperatures. The isotherms are 293 K (■), 313 K (▲), 333 K (●) and 353 K (◆). The lines represent a linear fit of the results for each isotherm, ($0.99998 \leq R^2 \leq 0.99999$).

The molar volume (V_m) of an ionic liquid was defined simply as the sum of the effective molar volumes (V^*) of the anion (a) and cation (c), Eq. (3):

$$V_m = V_c^* + V_a^* \quad (3)$$

The phosphonium ionic liquids presented in this study were divided into three groups, according to their structure,¹¹ and each group is represented as a function of temperature according to the following Eq. (4):

$$V_m(G_i) = A + B(T - 273.15) \quad (4)$$

V_m and is considered compose of contributions from $[P_{1111}]^+$, $-(CH_2)_n-$ and Cl^- , hence the over all molar volume is represented by Eq. (5):

$$V_m = \sum_{i=1}^3 V_m(G_i) \quad (5)$$

The parameters A and B can be obtained by a group contribution method according to Eq. xy. Eq. (6) and Eq. (7):

$$A = \sum_{i=1}^3 n_i a_i \quad (6)$$

$$B = \sum_{i=1}^3 n_i b_i \quad (7)$$

where n_i is the number of groups of type i , and the parameters a_i and b_i were previously estimated by Adamová and Gardas¹¹ on $[P_{8.8n}]Cl$ ionic liquids (Table 8).

Calculated molar volumes are in an excellent agreement with the corresponding experimental values (V_m)_{exp}, where:

$$(V_m)_{cal} = (1.0003 \pm 0.0002) \cdot (V_m)_{exp} \quad (R^2 = 0.9997 \text{ at } 95\% \text{ level of confidence}).$$

The relative deviations between the calculated and experimental data as a function of experimental molar volume for all data points are shown in Figure 22.

Table 8 Group contribution parameters associated with Eq. (5)¹¹ used to predict the effective molar volume of the phosphonium ionic liquids as a function of temperature from 293.15 to 363.15 K.

G_i	$a_i/cm^3 \text{ mol}^{-1}$	$b_i/cm^3 \text{ mol}^{-1} \text{ K}^{-1}$
$[P_{1111}]^+$	73.776	0.0840
$-CH_2-$	16.992	0.0082
Cl^-	24.391	0.0954

For 37 data points of five phosphonium ionic liquids, the overall AAD is 0.39% and MD is 0.80% (Table 9).

Table 9 AAD and MD observed for the molar volumes of studied phosphonium ionic liquids predicted using Eq.5.

n	AAD/%	MD/%
1	0.67	1.32
6	0.45	1.04
8	0.34	0.76
10	0.27	0.48
12	0.21	0.40

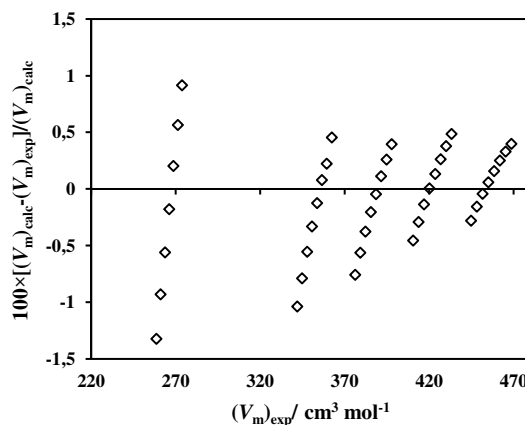


Fig. 22 The relative deviations between calculated, using an Eq.5 and the experimental molar volumes as a function of experimental molar volumes for the phosphonium ionic liquids, where $n = 1, 6, 8, 10$ and 12.

The deviations between calculated and experimental molar volumes are significantly higher, in contrast with $[P_{8.8n}]Cl$

compounds (AAD = 0.13% and MD = 0.28%),¹¹ which is most likely caused by much higher viscosity and mobility of [P_{444n}]Cl ionic liquids.

Thermal expansion coefficient

The evolution of the volumetric properties with temperature can be expressed by calculating the coefficient of thermal expansion, α_p , defined as, **Eq. (8)**:

$$\alpha_p = -[\partial \ln(\rho) / \partial T]_p = -a \quad (8)$$

Plots of $\ln(\rho)$ as a function of temperature, **Eq. (9)** (fitting equation parameters can be found in **Table 10**) for the studied ionic liquids have a *quasi*-linear fit, being almost parallel to each other with similar thermal expansion coefficients, α_p , given in **Table 10**.

$$\ln(\rho/cm^{-3}) = a(T/mK) + b \quad (9)$$

It has been already demonstrated by Tariq *et al.*,⁶⁶ that α_p may slightly decrease with increasing temperature in the low- T region and starts increasing (also mildly) in the high- T domain. This phenomenon was reported on the homologous series of [C_nmim][NTf₂] studied over extended range of temperatures, but here, with the limited temperature range no such deviation was found.

Additionally, we observe a dependency of α_p on chain length with calculated values lying between between $5.75 \times 10^{-4} \text{ K}^{-1}$ and $6.52 \times 10^{-4} \text{ K}^{-1}$, for [P₄₄₄₁]Cl and [P₄₄₄₁₂]Cl, respectively, see **Table 10**, as found in previous studies.^{10, 67, 68} The comparison of these data with related systems will be published elsewhere.⁶⁹

Table 10 Parameters of the linear fit, according to **Eq. (10)**

n	$a_p = -a/\text{K}^{-1} \times 10^3$	b
1	0.57467	0.136874
6	0.62610	0.115538
8	0.63546	0.108639
10	0.63870	0.102294
12	0.65240	0.099217
$\overline{\alpha_p} = (6.25464 \pm 0.29921) \times 10^{-4} \text{ K}^{-1}$		

Viscosity

The viscosity of ionic liquids is significantly influenced by the strength of both the van der Waals interaction forces and hydrogen bonding, but the chemical structure of the anion is known to have a strong influence, too. For example with 1,3-dialkylimidazolium halides, their high viscosity has been ascribed to an extensive three-dimensional hydrogen bonding network.⁷⁰ The viscosity has also been shown to be a colligative property, decreasing as a function of the number of moles of solute (either water or organic molecules).⁷¹

Viscosities of [P_{444n}]Cl, were measured at temperatures between 293.15 to 363.15 K (see **Table 11**). Viscosities decrease with increasing n , as illustrated in **Figure 23**. Moreover, the results indicate that their viscosity decreases markedly with temperature increase. For example, for [P₄₄₄₁₂]Cl, increasing the temperature by 10 K from (293.15 to 303.15) K decreases the viscosity by 57 % and increasing it by 70 K lowers it by 98 %, (see **Figure 24** and **25**).

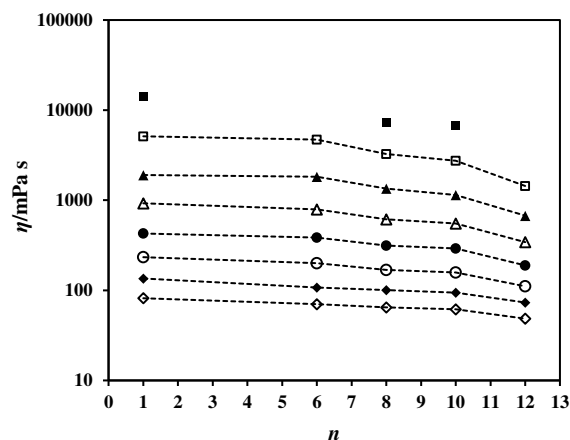


Fig. 23 Experimental viscosities of [P_{444n}]Cl ionic liquids at 293.15 K (■), 303.15 K (□), 313.15 K (▲), 323.15 K (△), 333.15 K (●), 343.15 K (○), 353.15 K (◆) and 363.15 K (◇) as a function of n .

The Vogel-Fulcher-Tammann (VFT) equation, **Eq. (10)**, which has been successfully used to model the temperature dependence of ionic liquid viscosity for many years,^{72,73} was applied to the data in **Table 10** and the fit is presented in **Figures 24** and **25**, and **Table 11**.

$$\eta = A + \exp(B/(T - T_0)) \quad (10)$$

where η is the viscosity in mPa s units, T is temperature (K), A , B and T_0 are adjustable parameters.

The ratio of parameters B and T_0 , B/T_0 , is also known as the Angell strength parameter. The parameters A , B and T_0 were optimised by minimising the following objective function (OF), **Eq. (11)**:

$$OF = \sum_{i=1}^{N_p} \left[A + \frac{B}{(T - T_0)} - \ln(\rho_{exp}) \right]^2 / N_p \quad (11)$$

The AAD and MD values are presented in **Table 12**. The values of the parameter T_0 are between 105 to 160 K.

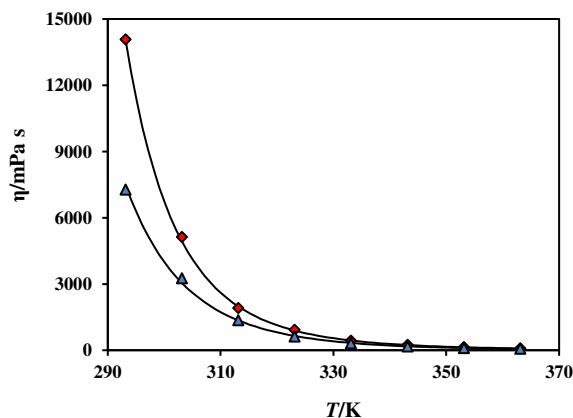


Fig. 24 Experimental viscosities as a function of temperature: [P₄₄₄₁]Cl (◆) and [P₄₄₄₈]Cl (▲), fitted into VFT equation, **Eq. 11**. Fits are represented by solid lines.

Table 11 Experimental viscosities (η_{exp}) of phosphonium chlorides as a function of temperature at atmospheric pressure. Dev^a represents standard deviation of ten measurements performed with each sample at each given temperature. Calculated viscosities (η_{cal}) were obtained from VFT equation and AAD represents agreement between calculated and experimental data.

<i>T</i> / K	η_{exp} / mPa s	Dev ^a / %	η_{cal} / mPa s	AAD ^b / %
[P₄₄₄₁]Cl				
293.15	14065.77	2.14	14061.85	0.03
303.15	5123.50	2.03	4925.49	3.86
313.15	1904.94	0.91	1978.55	3.86
323.15	924.50	0.97	888.78	3.86
333.15	427.00	0.92	437.91	2.56
343.15	233.20	1.33	233.10	0.04
353.15	135.30	2.27	132.45	2.11
363.15	81.70	3.01	79.57	2.61
[P₄₄₄₆]Cl				
293.15	-	-	-	-
303.15	4710.30	1.27	4631.80	1.67
313.15	1815.00	1.63	1813.06	0.11
323.15	788.61	1.76	795.38	0.86
333.15	386.23	3.09	383.47	0.72
343.15	200.08	3.06	200.08	0.00
353.15	107.30	1.85	111.62	4.03
363.15	70.14	3.51	65.92	6.01
[P₄₄₄₈]Cl				
293.15	7269.56	1.19	7486.04	2.98
303.15	3249.99	1.19	3043.91	6.34
313.15	1349.48	1.64	1349.48	0.00
323.15	614.12	1.99	644.60	4.96
333.15	313.86	3.03	328.51	4.67
343.15	168.23	2.78	177.17	5.31
353.15	100.44	3.31	100.42	0.02
363.15	64.85	2.69	59.48	8.27
[P₄₄₄₁₀]Cl				
293.15	6802.34	4.32	6765.87	0.54
303.15	2742.77	6.41	2633.04	4.00
313.15	1143.27	2.79	1149.58	0.55
323.15	552.03	4.95	552.03	0.00
333.15	291.05	4.43	287.07	1.37
343.15	157.37	1.62	159.69	1.47
353.15	94.09	2.81	94.08	0.00
363.15	61.43	2.99	58.24	5.20
[P₄₄₄₁₂]Cl				
293.15	-	-	-	-
303.15	1439.04	4.15	1429.76	0.64
313.15	674.30	4.21	674.94	0.09
323.15	342.50	1.91	345.98	1.02
333.15	189.05	2.49	190.13	0.57
343.15	110.86	3.93	110.86	0.00
353.15	73.25	2.84	68.01	7.15
363.15	48.50	2.30	43.60	10.12

Table 12 Correlation parameters of the VFT equation, Eq.11, with the deviations of the fit for the viscosity of [P_{444*n*}]Cl as a function of temperature, determined from experimental values in the range $T = 293.15$ to $T = 363.15$ K.}

<i>n</i>	<i>A</i>	<i>B</i> /K	<i>T</i> ₀ /K	<i>B</i> / <i>T</i> ₀	AAD/ %	MD/ %
1	-12.38	2000.00	159.98	12.50	2.4	3.9
6	-12.97	2096.43	158.56	13.22	1.9	6.0
8	-15.82	3354.43	105.02	31.94	4.1	8.3
10	-12.63	2096.27	149.03	14.07	1.6	5.2
12	-12.57	2095.86	140.98	14.87	2.8	10.1

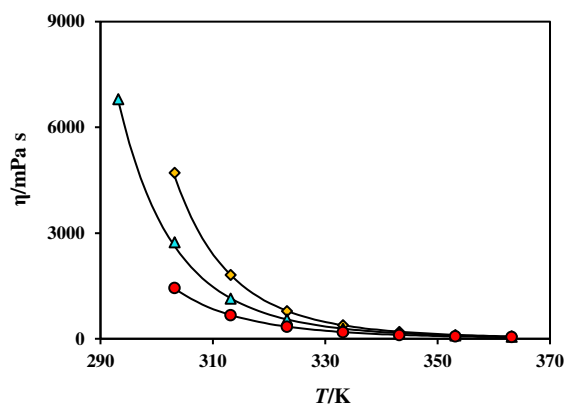


Fig. 25 Experimental viscosities as a function of temperature: [P₄₄₄₆]Cl (♦), [P₄₄₄₁₀]Cl (▲) and [P₄₄₄₁₂]Cl (●) fitted into VFT equation, Eq. 11. Fits are represented by solid lines.

The experimental viscosity data are plotted against data calculated from Eq. 10 with the fitted data in Table 12, represented in Figure 26; excellent agreement is revealed by the fit in Eq. 12

$$\ln(\eta_{\text{cal}}) = (0.9974 \pm 0.0032)\ln(\eta_{\text{exp}}) \quad (12)$$

$(R^2 = 0.9995 \text{ at a } 95\% \text{ of confidence})$

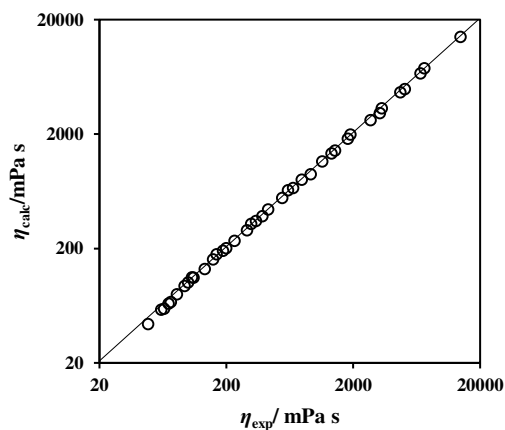


Fig. 26 Linear relationship between η_{exp} and η_{cal} using VFT equation, Eq. 11

Relative deviations between the calculated and the experimental viscosity data are presented in Table 11 along with experimental viscosity values over a wide range of temperatures (293.15-363.15 K).

Conclusions

The series of ionic liquids presented here are all hydrophilic and present an interesting contrast with hydrophobic [P_{888*n*}]Cl. Of particular interest is the crystal structure of [P₄₄₄₃]Cl which allows a comparison to be made with its tetraalkylammonium analogues. The combination of electrostatic attraction, hydrogen bonding and steric repulsion leads to a delicate balance between these forces. The shorter nitrogen carbon bonds of the tetraalkylammonium structures leads to a sterically compressed “shell” of methylene groups, which prevent the close approach of the chloride ion that might have been expected from a facile}

consideration of the size of the nitrogen atom. In contrast, the larger phosphorus atom and longer phosphorus carbon bonds leads to a more open structure, permitting the chloride ion to approach closer to the centre of the cation. Although counter-intuitive, the observed structures lead to an intellectually satisfying interpretation of some of the contrasting properties of tetraalkylammonium and tetraalkylphosphonium ionic liquids.

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